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INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

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(21) International Application Number: PCT/US95/06110 (22) International Filing Date: 15 May 1995 (15.05.95) (30) Priority Data: 08/246,962 20 May 1994 (20.05.94) US 08/440,450 12 May 1995 (12.05.95) US (71) Applicant: MINNESOTA MINING AND MANUFACTURING COMPANY [US/US]; 3M Center, P.O. Box 33427, Saint Paul, MN 55133-3427 (US). (72) Inventors: MOORE, George, G., I.; P.O. Box 33427, Saint Paul, MN 55133-3427 (US). FLYNN, Richard, M.; P.O. Box 33427, Saint Paul, MN 55133-3427 (US). GUERRA, Miguel, A.; P.O. Box 33427, Saint Paul, MN 55133-3427 (US). OWENS, John, G.; P.O. Box 33427, Saint Paul, MN 55133-3427 (US). (74) Agents: WEISS, Lucy, C. et al.; Minnesota Mining and Manufacturing Company, Office of Intellectual Property Counsel, P.O. Box 33427, Saint Paul, MN 55133-3427 (US).		(81) Designated States: CA, CN, JP, RU, European patent (AT, BE, CH, DE, DK, ES, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE). Published <i>With international search report. Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</i>
(54) Title: OMEGA-HYDROFLUOROALKYL ETHERS, PRECURSOR CARBOXYLIC ACIDS AND DERIVATIVES THEREOF, AND THEIR PREPARATION AND APPLICATION		
(57) Abstract Normally liquid, omega-hydrofluoroalkyl ether compounds (and selected mixtures thereof) have a saturated perfluoroalkyl chain of carbon atoms interrupted by one or more ether oxygen atoms. The compounds can be prepared, e.g., by decarboxylation of the corresponding fluoroalkyl ether carboxylic acids and are useful, e.g., in cleaning and drying applications.		

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5 OMEGA-HYDROFLUOROALKYL ETHERS, PRECURSOR CARBOXYLIC
ACIDS AND DERIVATIVES THEREOF, AND THEIR
PREPARATION AND APPLICATION

10 This invention relates to omega-hydrofluoroalkyl
ethers and their preparation and application. In
another aspect, this invention relates to perfluoro-
(alkoxyalkanoic) acids and derivatives thereof and
their preparation. In another aspect, it relates to
the preparation of perfluoro(alkoxyalkanoic) acids by
direct fluorination of their hydrocarbon alkanoic acid
15 or ester analogs and to the preparation of omega-
hydrofluoroalkyl ethers, for example, by
decarboxylation of said acids or their alkyl esters.
In another aspect, this invention relates to uses of
perfluoro(alkoxyalkanoic) acids and derivatives
20 thereof.

Because of a steady flow of bad news about the
damaged stratospheric ozone layer, the deadlines for
the end to industrialized countries' production of
chlorofluorocarbons ("CFCs") and other ozone-depleting
25 chemicals were accelerated by countries who are parties
to the Montreal Protocol on Substances That Deplete the
Ozone Layer - see Zurer, P.S., "Looming Ban on
Production of CFCs, Halons Spurs Switch to
Substitutes," November 15, 1993, Chemical & Engineering
30 News, p. 12.

Work is under way to replace CFCs and halons, such
as CCl_2F_2 , CCl_3F , CF_3Br , and $\text{CCl}_2\text{FCClF}_2$, with substitute
or alternative compounds and technologies. A number of
hydrofluorocarbons ("HFCs"), e.g., CH_2FCF_3 ("HFC-134a"),
35 are being used or have been proposed as CFC substitutes
(and HFC-134a has been characterized as being more
"ozone friendly" - see U.S. Patent No. 5,118,494

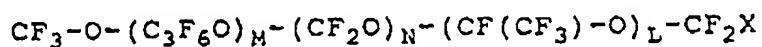
(Schultz et al.)). Hydrochlorofluorocarbons ("HCFCs"), such as $\text{CH}_3\text{CCl}_2\text{F}$ ("HCFC-141b"), as the C&EN article, supra, points out, are CFC substitutes, but although they are not nearly as damaging, these substitutes do carry ozone-depleting chlorine into the stratosphere. Another proposed substitute is the simple omega-hydrodifluoromethyl perfluoromethyl ether, $\text{CF}_3\text{OCF}_2\text{H}$ -- See J.L Adcock et.al., "Fluorinated Ethers -- A new Family of Halons," 1991 CFC Conference Proceedings (1991). Another hydro-fluoroalkyl ether (or ether hydride), $\text{F}[\text{CF}(\text{CF}_3)\text{CF}_2\text{O}]_4\text{CFHCF}_3$, made by decarboxylation of the fluorinated 2-alkoxypropionic acid salt, has been tested as a blood emulsion -- see Chem. Pharm. Bull. 33, 1221 (1985).

U.S. Pat. No. 4,173,654 (Scherer) states that fluorocarbons due to their inertness have found use as electronic coolant or leak testing fluids, and other compounds having good solubility for oxygen have been investigated as artificial blood substitutes. This patent describes certain fluorocarbon "hybrid" materials with metabolically active hydrocarbon moieties, such as, inter alia, $-\text{CH}_2-(\text{CH}_2)_m-\text{H}$. U.S. Pat. No. 4,686,024 (Scherer et al.), which describes certain perfluorocyclic ethers, states that various perfluoro chemicals are disclosed in patents as being suitable as oxygen and carbon dioxide carriers. And International Application published as WO 93/11868 (Kaufman et al.) describes certain chlorofluorochemicals and emulsions thereof as useful in various oxygen transport applications, e.g., as oxygen transfer agents or "artificial bloods."

There are a number of other patents describing various fluorocarbon ethers or polyethers. U.S. Patent No. 3,342,875 (Selman et al.) describes certain "hydrogen modified fluorocarbon ethers" (or "hydrogen capped polyethers") made, inter alia, by pyrolysis of a hydrogen-containing derivative of an ether, such as

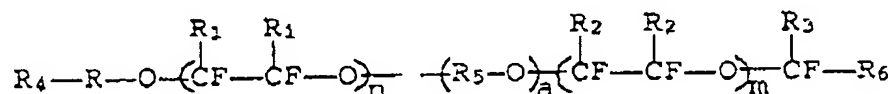
the fluorocarbon ether acid or the ammonium salt, which ether is obtained by the polymerization of fluorocarbon epoxides. British Patent Specification 1,194,431 (Montecatini Edison S.P.A.) describes certain

5 perfluorinated ethers and polyether derivatives having the general formula



10 where, inter alia, each subscript M, N, and L is zero or a whole number from 1 to 99, and X is a hydrogen atom or -COOMe wherein Me is an equivalent of an alkali or alkaline earth metal, an examples of which is pentafluorodimethyl ether, $\text{CF}_3\text{-O-CF}_2\text{H}$.

15 U.S. Patent No. 3,597,359 (Smith) describes certain perfluoroalkylene ether-containing compound represented by the formula



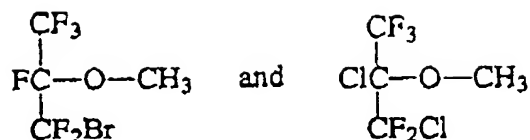
20 wherein, inter alia, R is alkylene, alkoxyalkylene, or perfluoroalkylene, R_1 is fluorine or trifluoromethyl provided not more than one R_1 is trifluoromethyl, R_2 is fluorine or trifluoromethyl provided not more than one R_2 is trifluoromethyl, R_3 is fluorine or

25 trifluoromethyl, R_4 is hydrogen or halogen provided that when R is alkylene or alkoxyalkylene R_4 is hydrogen, R_5 is perfluoroalkylene having at least 2 carbon atoms, R_6 is, inter alia hydrogen, trifluoromethyl or perfluoroethyl, a is zero or 1, n

30 and m are whole numbers of 0 to 50, and $n + m$ is 1 to 50.

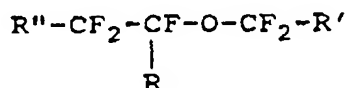
U.S. Patent No. 3,962,460 (Croix et al.) describes aliphatic ethers, including those of the formulas

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10 International Patent Application WO 90/01901 (Long)
describes certain perfluorocarbon hydrides, such as
perfluorooctyl hydride, used in emulsions for carrying
oxygen to the tissues of an animal body. European
Patent Application Publication No. 0 482 938 A1
15 (Chambers et al.) describes fluorinated ethers of the
formula

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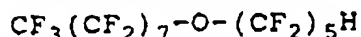
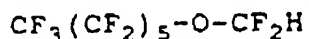
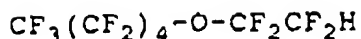


wherein R is hydrogen, fluorine, or alkyl or
fluoroalkyl of 1-6 carbon atoms, R' is hydrogen or
alkyl or fluoroalkyl of 1 to 6 carbon atoms, and R'' is
25 fluorine or alkyl or fluoroalkyl of 1 to 6 carbon
atoms.

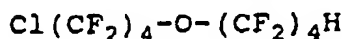
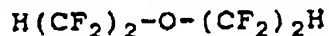
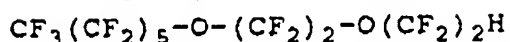
Other patents describing one or more various
fluoroalkoxyalkanoic acids and esters or other
derivatives thereof and their preparation are U.S.
30 Patent Nos. 2,713,593 (Brice et al.), 3,214,478
(Milian, Jr.), 3,393,228 (Braun), 4,118,421 (Martini),
4,357,282 (Anderson et al.), 4,729,856 (Bernonge),
4,847,427 (Nappa), 4,940,814 (Schwertfeger), 4,973,716
(Calini et al.), 5,053,536 (Bierschenk et al.)
35 5,093,432 (Bierschenk et al.), and 5,118,494 (Schultz
et al.) and PCT International Applications Pub. Nos. WO
90/03357 (Moore et al.) and WO 90/06296 (Costello et
al.). The aforementioned Brice et al. patent describes
fluorocarbons acids made by electrochemical
40 fluorination including an acid having a boiling point
of 225°C and said to be n-C₈F₁₇OC₂F₄CO₂H. The
aforementioned Nappa, Bierschenk et al., Moore et al.,

and Costello et al. publications describe the preparation of the fluorinated compounds by direct fluorination of hydrocarbon analog precursors.

In one aspect, this invention provides a normally liquid (i.e., liquid under ambient conditions of temperature and pressure) fluoroalkyl ether compound or a normally liquid composition consisting or consisting essentially of a selected mixture of such compounds, said compound having a saturated perfluoroaliphatic chain of carbon atoms (e.g., 4 to 30) interrupted by one or a plurality (e.g., 2 to 8) of ether (or catenary, i.e., in-chain) oxygen atoms. The chain carbon atom at one end (hereafter called the proximal end) of the chain is bonded to a hydrogen atom (i.e., an omega-hydro substituent, or primary hydrogen atom) and two fluorine atoms, said proximal carbon atom being the carbon atom of a difluoromethyl group or moiety, $-CF_2H$, which is directly bonded to another chain carbon atom, such as that of perfluoroalkylene chain segment, $-C_nF_{2n}$, or to a said ether-oxygen. The carbon atom at the other end of the chain (the distal end) is part of a distal group selected from the group consisting of a difluoromethyl, a difluorochloromethyl, $-CF_2Cl$, a perfluoroalkyl substituted with a saturated alicyclic moiety, e.g., $c-C_6F_{11}-$, a straight-chain perfluoroalkyl, and a branched chain perfluoroalkyl. In a said compound where said proximal end of the chain terminates in a difluoromethyl group bonded to an ether-oxygen atom, then said straight-chain perfluoroalkyl has at least 6 chain carbon atoms, e.g., 6 to 16 chain carbon atoms, and said branched-chain perfluoroalkyl has at least 4 carbon atoms, e.g., 4 to 16 carbon atoms. Examples of such omega-hydro fluoroalkyl ether compounds are:



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10 If a said "selected mixture," i.e., a predetermined mixture of selected omega-hydrofluoroalkyl ether compounds, is desired for a particular use, a said composition of this invention can be made consisting or consisting essentially of a mixture of two or more of said compounds each having a desired discrete, non-
15 random molecular weight, the selected compounds preferably being those having complementary properties, e.g., for imparting improved stability to emulsions where they are incorporated as oxygen carriers in medical applications.

20 The term "perfluoro," such as in the case of "perfluoroaliphatic," "perfluoroalkylene," or "perfluoroalkyl," means that except as may be otherwise indicated there are no carbon-bonded hydrogen atoms replaceable with fluorine nor any unsaturation.

25 Omega-hydrofluoroalkyl ethers of this invention are hydrophobic and less oleophobic than the perfluoroalkyl ether analogs, chemically inert, thermally stable, water insoluble, and normally liquid (e.g., at 20°C), and they can be made in accordance
30 with this invention in high yield, high purity, and with a wide range of molecular weights. The covalent bond between the omega-hydrogen and terminal carbon, i.e., the C-H bond, is generally degradable by atmospheric photo-oxidation, thus making the omega-hydrofluoroalkyl ethers environmentally acceptable or
35 compatible. The omega-hydrofluoroalkyl ether compounds, or the normally liquid composition

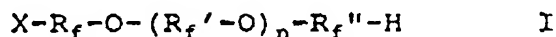
consisting or consisting essentially thereof, can be used in applications where the aforementioned CFCs, HCFCs or halons have been used, for example, as solvents for precision or metal cleaning of electronic articles such as disks or circuit boards, heat transfer agents, coolants in refrigerator or freezer compressors or air conditioners, blowing agents or cell size regulators in making polyurethane foam insulation, or chemical fire extinguishing agents in streaming applications, total flooding, explosion suppression and inertion, and as carrier solvents for highly fluorinated polyethers used as lubricants for magnetic recording media. Another field of utility for the omega-hydrofluoroalkyl ethers is in emulsions useful in various medical and oxygen transport applications, for example, artificial or synthetic bloods.

The above-described omega-hydrofluoroalkyl ethers of this invention can be prepared by decarboxylation of the corresponding precursor fluoroalkyl ether carboxylic acids and salts thereof or, preferably, the saponifiable alkyl esters thereof. Alternatively, the omega-hydrofluoroalkyl ethers can be prepared by reduction of the corresponding omega-chlorofluoroalkyl ethers (e.g., those described in WO 93/11868, supra). The perfluoroalkyl ether carboxylic acids (and esters) themselves -- some of which are believed novel compounds and they and their preparation are other aspects of this invention -- can be prepared by direct fluorination of their corresponding hydrocarbon analogs. The omega-hydrofluoroalkyl ethers are essentially pure fluorinated compounds and are useful as such or in the form of a normally liquid composition consisting or consisting essentially of a selected mixture of such compounds. The precursor perfluoroalkyl ether carboxylic acid and ester compounds, like the above-described omega-hydrofluoroalkyl compounds of this invention, have a

saturated perfluoroaliphatic chain of a plurality of carbon atoms, said chain likewise being interrupted by one or a plurality of ether oxygen atoms, the proximal end of the chain being connected to a carboxyl group or alkyl ester thereof. This carboxyl group (or salts thereof or its saponifiable alkyl ester) can be decarboxylated, as mentioned above, and thereby replaced by the aforementioned omega-hydro substituent of the resulting omega-hydroalkyl ether of this invention.

The aforementioned novel perfluoroalkyl ether acids and esters can also be converted into various other derivatives, such as their ammonium salts, which have utility as surface active agents useful in modifying the surface tension or interfacial tension of liquids. These compounds are more soluble in aqueous media and other organic solvents than are the corresponding perfluoroalkanoic acid derivatives, and this enhances their utility as surface-active agents. The compounds can conveniently be prepared by direct fluorination of the corresponding hydrocarbon ether acids, or derivatives such as an ester, in high yields as single molecular species.

A class of the normally liquid, omega-hydrofluoroalkyl ether compounds of this invention can be represented by the general formula:



wherein:

H is a primary hydrogen atom;

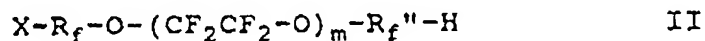
X is a fluorine atom, a primary hydrogen atom, or a primary chlorine atom bonded to a difluoromethylene (of R_f);

n is a integer of 0 to 7, preferably 0 to 3;

R_f , R_f' , and R_f'' are the same or different perfluoroalkylene (linear or branched) groups, e.g., $-\text{CF}_2\text{CF}_2-$, which are unsubstituted or substituted with a perfluoro organo group which can contain ether oxygen, for example, R_f can be $-\text{CF}_2\text{CF}(R_f'')\text{CF}_2-$ or $-\text{R}_f''\text{CF}_2-$ where R_f'' is a saturated perfluoroalicyclic group having 4 to 6 ring carbon atoms, such as perfluorocyclohexyl or perfluorocyclohexylene;

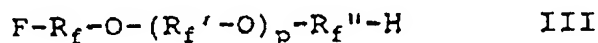
with the proviso that when X is H or Cl, R_f has 1 to 18, preferably 2 to 18, chain carbon atoms, R_f' has 1 to 12, preferably 2 to 12, chain carbon atoms, and R_f'' has 2 to 12 chain carbon atoms; and with the further proviso that when X is F, then R_f has at least 4, preferably 4 to 18, chain carbon atoms, R_f' has 1 or more, preferably 1 to 12, more preferably 2 to 12, chain carbon atoms, and R_f'' has 2 or more, preferably 2 to 12, chain carbon atoms.

A subclass of polyether compounds within the scope of general formula I is represented by the general formula:



where m is an integer of 0 to 7, and H, X, R_f , and R_f'' are as defined for formula I.

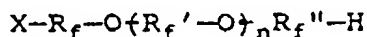
Another subclass of compounds within the scope of general formula I is represented by the general formula:



where p is an integer of 0 to 2 and H, R_f , R_f' , and R_f'' are as defined for formula I, except R_f has 4 to 12

chain carbon atoms, R_f' has 1 to 12 chain carbon atoms, and R_f'' has 2 to 12 chain carbon atoms.

5 Another class of the normally liquid, omega-hydrofluoroalkyl ether compounds of the invention can be represented by the general formula:



10 wherein:

H is a primary hydrogen atom;

X is a fluorine atom, a primary hydrogen atom, or a primary chlorine atom;

n is an integer of 0 to 7; and

15 R_f , R_f' , and R_f'' are independently selected from the group consisting of linear or branched, unsubstituted perfluoroalkylene groups; linear or branched, perfluoroalkyl- or perfluorocycloalkyl-substituted
20 perfluoroalkylene groups; and linear or branched perfluoroalkylene groups substituted with an ether oxygen-containing moiety;

with the proviso that when X is H or Cl, R_f has 1 to 18 chain carbon atoms and each of R_f' and R_f''

25 independently has 1 to 12 chain carbon atoms;

and with the further proviso that when X is F, then R_f has at least 4 chain carbon atoms and each of R_f' and R_f'' independently has 1 or more chain carbon atoms;

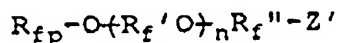
30 and with the still further proviso that when n is zero, then R_f is a perfluorocycloalkyl-substituted perfluoroalkylene group.

A list of representative examples of the omega-hydrofluoroalkyl ether compounds of this invention is
35 as follows.

Table A

1. $\text{CF}_3(\text{CF}_2)_5\text{-O-CF}_2\text{H}$
2. $\text{CF}_3(\text{CF}_2)_3\text{-O-(CF}_2)_2\text{H}$
3. $\text{C-C}_6\text{F}_{11}\text{CF}_2\text{-O-(CF}_2)_2\text{H}$
4. $\text{CF}_3(\text{CF}_2)_3\text{-O-CF}_2\text{C(CF}_3)_2\text{CF}_2\text{H}$
5. $(\text{CF}_3)_2\text{CFCF}_2\text{-O-CF}_2\text{H}$
6. $\text{CF}_3(\text{CF}_2)_4\text{-O-(CF}_2)_5\text{H}$
7. $\text{CF}_3(\text{CF}_2)_6\text{-O-CF}_2\text{H}$
8. $\text{CF}_3(\text{CF}_2)_5\text{-O-(CF}_2)_2\text{H}$
9. $\text{CF}_3(\text{CF}_2)_5\text{-O-(CF}_2)_3\text{H}$
10. $\text{CF}_3(\text{CF}_2)_6\text{-O-(CF}_2)_2\text{H}$
11. $\text{CF}_3(\text{CF}_2)_7\text{-O-CF}_2\text{H}$
12. $\text{CF}_3(\text{CF}_2)_7\text{-O-(CF}_2)_5\text{H}$
13. $\text{CF}_3(\text{CF}_2)_7\text{-O-(CF}_2)_6\text{H}$
14. $\text{CF}_3(\text{CF}_2)_5\text{-O-(CF}_2)_2\text{-O-CF}_2\text{H}$
15. $\text{CF}_3(\text{CF}_2)_5\text{-O-(CF}_2)_2\text{-O-(CF}_2)_2\text{H}$
16. $\text{H-(CF}_2)_2\text{-O-(CF}_2)_2\text{H}$
17. $\text{H-(CF}_2)_4\text{-O-(CF}_2)_4\text{H}$
18. $\text{H-(CF}_2)_2\text{-O-(CF}_2)_2\text{-O-(CF}_2)_2\text{H}$
19. $\text{H-CF}_2\text{-O-CF}_2\text{C(CF}_3)_2\text{CF}_2\text{-O-CF}_2\text{H}$
20. $\text{Cl(CF}_2)_4\text{-O-(CF}_2)_4\text{H}$
21. $\text{H(CF}_2)_2\text{OCF}_2\text{C(CF}_3)_2\text{CF}_2\text{O(CF}_2)_2\text{H}$
22. $\text{C}_8\text{F}_{17}\text{OCF}_2\text{OC}_3\text{F}_6\text{H}$
23. $(\text{CF}_3)_3\text{COC}_2\text{F}_4\text{OCF}_2\text{OC}_2\text{F}_4\text{OCF}_2\text{H}$

As mentioned above, the omega-hydrofluoroalkyl ether compounds or compositions of this invention can be made by decarboxylation of their corresponding precursor perfluoroalkyl ether carboxylic acids, hydrolyzable carboxylic acid derivatives, or hydrolyzable precursors thereto (some of which are believed novel). A class of such precursor compounds can be represented by the general formula:



IV

wherein

R_{fp} is $ROC(O)R_f$ or $F-R_f$, R_f being a perfluoroalkylene group as defined for formula I;

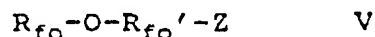
5 R_f' and R_f'' are also perfluoroalkylene groups as defined for formula I;

n is also as defined for formula I; and

10 Z' is a CO_2H , CO_2R , COF , $COCl$, $CONR^1R^2$, or $-CF_2OC(O)R_f$, where R is selected from the group consisting of hydrogen, alkyl (such as a lower alkyl group of 1 to 6 carbon atoms), cycloalkyl, fluoroalkyl, and aryl, and where R^1 and R^2 are independently selected from the group consisting of hydrogen, alkyl, 15 cycloalkyl, and heteroatom-containing cycloalkyl.

In the decarboxylation of the compounds of formula IV, the moiety Z' is replaced by a hydrogen atom.

20 Subclasses of said ether acids and derivatives thereof, which have other utilities in addition to their use as precursors of the omega-hydro ether compounds of this invention, for example, as surface active agents (or surfactants), as mentioned above, and 25 which are believed novel, can be represented by the general formulas V, VI, VII, VIII and IX below,



30 wherein:

R_{fo} is a perfluoroalkyl group (linear or branched) having, for example, 1 to 18 carbon atoms, preferably 1 to 12 carbon atoms,

35 R_{fo}' is a perfluoroalkylene group (linear or branched) having, for example, 2 to 11 carbon atoms, at least one of R_{fo} and R_{fo}' having at least 8 chain carbon atoms; and

Z is $-\text{COOH}$, $-\text{COOM}_{1/v}$, $-\text{COONH}_4$, $-\text{COOR}$, $-\text{CH}_2\text{OH}$, $-\text{COF}$,
 $-\text{COCl}$, $-\text{CR}$, $-\text{CONRR}$, $-\text{CH}_2\text{NH}_2$, $-\text{CH}_2\text{NCO}$, $-\text{CN}$,
 $-\text{CH}_2\text{OSO}_2\text{R}$, $-\text{CH}_2\text{OCOR}$, $-\text{CH}_2\text{OCOCR}=\text{CH}_2$,
 $-\text{CONH}(\text{CH}_2)_m\text{Si}(\text{OR})_3$, or $-\text{CH}_2\text{O}(\text{CH}_2)_m\text{Si}(\text{OR})_3$,
 where M is an ammonium radical or a metal
 atom having a valence "v" of 1 to 4, such as
 Na, K, Ti, or Al, and each R is independently
 an alkyl (e.g., with 1 to 14 carbon atoms) or
 cycloalkyl, which groups can be partially or
 fully fluorinated, or an aryl (e.g., with 6
 to 10 ring-carbon atoms), any of which groups
 can contain heteroatom(s), and m is an
 integer of 1 to about 11.

15 $\text{R}_{\text{fq}}\{\text{O}-\text{CF}_2\text{CF}_2\}_a\text{OCF}_2-\text{Z}$ VI

wherein:

R_{fq} is a perfluoroalkyl group (linear or branched)
 having from about 6 to about 18 carbon atoms,
 preferably 6 to 12 carbon atoms,

20 subscript a is an integer of at least 2,
 preferably 3 to 7, but when a is 2, then R_{fq}
 has at least about 8 carbon atoms; and
 Z is as defined for formula V.

25 $\text{R}_{\text{fr}}-\text{O}-\text{CF}_2-\text{O}-\text{R}_{\text{fr}}'-\text{Z}$ VII

wherein:

R_{fr} is a perfluoroalkyl group (linear or branched)
 having, for example, 2 to 18 carbon atoms,
 preferably 4 to 12 carbon atoms;

30 R_{fr}' is a perfluoroalkylene group (linear or
 branched) having, for example, 1 to 11 carbon
 atoms and preferably 1 to 5 carbon atoms; and
 Z is as defined for formula V; and the sum of the
 number of carbon atoms in the groups R_{fr} and
 35 R_{fr}' is at least about 7.

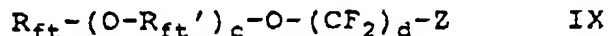
$\text{R}_{\text{fb}}-\text{O}\{\text{CF}_2\}_b-\text{Z}$ VIII

wherein:

R_{fs} is a perfluoroalkyl group (linear or branched) having, for example, 1 to 18 carbon atoms, preferably 1 to 12 carbon atoms;

5 b is an integer of at least 3, preferably 3 to 11; and

z is as defined for formula V.



10 wherein:

R_{ft} is a perfluoroalkyl group (linear or branched) having, for example, 1 to 18 carbon atoms, preferably 1 to 12 carbon atoms;

15 R_{ft}' is a perfluoroalkylene group (linear or branched) having, for example, 1 to 11 carbon atoms, preferably 2 to 4 carbon atoms;

c is an integer of at least 1, preferably 1 to 4;

d is an integer of 3 or greater, preferably 3 to 9; and

20 z is as defined for formula V.

The carboxylic acids of formulas V to IX are useful intermediates for the preparation of many of the other derivatives of formulas V to IX. These derivatives include nonfunctional or functional derivatives such as, for example, carboxylic acids, salts, esters, amides, nitriles, alcohols, acrylates, and vinyl ethers. Various patents describe processes for the preparation of a host of functional derivatives of oxyperfluoroalkylene compounds, i.e., perfluoropolyethers, e.g., see U.S. Pat. Nos. 3,250,808 (Mitsch et al.) and 4,094,911 (Moore et al.). These derivatives have utility for various applications, such as surfactants, elastomers, coatings, lubricants, substances used in the preparation of liquid crystal materials such as those in U.S. 5,262,082 (Janulis et al.), and in the treatment of fibrous substrates to

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30
35

impart oil and water repellency thereto. The ammonium salts of the carboxylic acid derivatives are particularly useful as surfactants.

5 The carboxylic acid compounds of formula V are normally solid. The carboxylic acid compounds of formulas VI, VII, VIII and IX generally are normally liquid and normally liquid compositions can be made up which consist or consist essentially of selected mixtures of such compounds.

10 A list of representative examples of fluoroalkylether acids (or derivatives) which can be utilized to prepare omega-hydrofluoroalkyl ethers of this invention is as follows:

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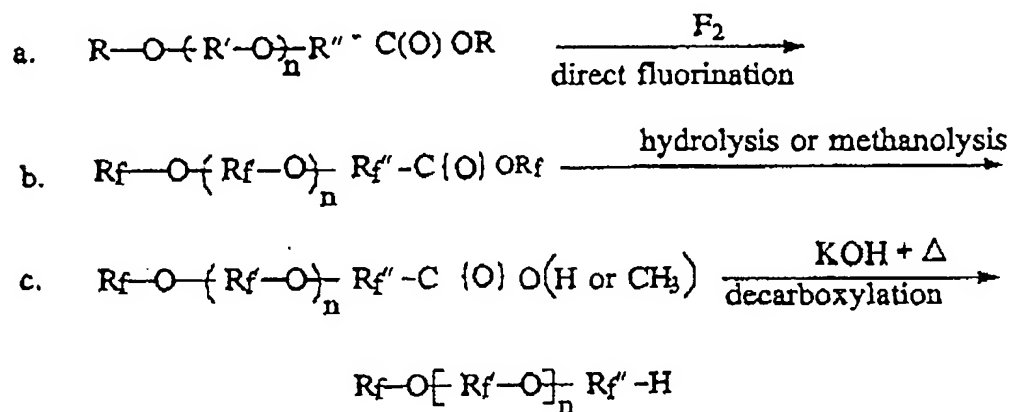
Table B

1. $\text{CF}_3(\text{CF}_2)_7\text{-O-CF}_2\text{CO}_2\text{H}$
2. $\text{CF}_3(\text{CF}_2)_{11}\text{-O-CF}_2\text{CO}_2\text{H}$
3. $\text{CF}_3(\text{CF}_2)_6\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$
4. $\text{CF}_3(\text{CF}_2)_4\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$
- 20 5. $\text{CF}_3(\text{CF}_2)_5\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$
6. $\text{CF}_3(\text{CF}_2)_8\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$
7. $\text{CF}_3(\text{CF}_2)_7\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$
8. $\text{CF}_3(\text{CF}_2)_9\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$
9. $\text{CF}_3(\text{CF}_2)_{11}\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$
- 25 10. $\text{CF}_3(\text{CF}_2)_5\text{-OC}_2\text{F}_4\text{O-C}_2\text{F}_4\text{CO}_2\text{H}$
11. $\text{C}_8\text{F}_{17}\text{-O-(CF}_2)_5\text{CO}_2\text{H}$
12. $\text{C}_{10}\text{F}_{21}\text{-O-(CF}_2)_5\text{CO}_2\text{H}$
13. $\text{CF}_3\text{-O-(CF}_2)_7\text{CO}_2\text{H}$
14. $\text{C}_2\text{F}_5\text{-O-(CF}_2)_7\text{CO}_2\text{H}$
- 30 15. $\text{C}_3\text{F}_7\text{-O-(CF}_2)_7\text{CO}_2\text{H}$
16. $\text{CF}_3\text{-O-(CF}_2)_9\text{CO}_2\text{H}$
17. $\text{CF}_3\text{-O-(CF}_2)_{10}\text{CO}_2\text{H}$
18. $\text{CF}_3(\text{CF}_2)_5\text{-O-C}_2\text{F}_4\text{-O-C}_2\text{F}_4\text{-O-C}_2\text{F}_4\text{-O-CF}_2\text{CO}_2\text{H}$
- 35 19. $\text{CF}_3(\text{CF}_2)_7\text{-O-C}_2\text{F}_4\text{-O-C}_2\text{F}_4\text{-O-C}_2\text{F}_4\text{-O-CF}_2\text{CO}_2\text{H}$

20. $\text{CF}_3(\text{CF}_2)_9\text{-O-C}_2\text{F}_4\text{-O-C}_2\text{F}_4\text{-O-C}_2\text{F}_4\text{-O-}$
 $\text{CF}_2\text{CO}_2\text{H}$
21. $\text{CF}_3(\text{CF}_2)_{11}\text{-O-C}_2\text{F}_4\text{-O-C}_2\text{F}_4\text{-O-C}_2\text{F}_4\text{-O-}$
 $\text{CF}_2\text{CO}_2\text{H}$
- 5 22. $\text{CF}_3(\text{CF}_2)_{11}\text{-(OC}_2\text{F}_4\text{)}_{1-5}\text{-O-CF}_2\text{CO}_2\text{H}$ from
Brijtm30 acetate
23. $\text{C}_6\text{F}_{13}\text{OCF}_2\text{O}(\text{CF}_2)_5\text{CO}_2\text{H}$
24. $\text{CF}_3(\text{CF}_2)_7\text{-O-CF}_2\text{-O-CF}_2\text{CO}_2\text{H}$
25. $\text{CF}_3(\text{CF}_2)_7\text{-O-CF}_2\text{-O-C}_3\text{F}_6\text{CO}_2\text{H}$
- 10 26. $(\text{CF}_3)_3\text{COC}_2\text{F}_4\text{OCF}_2\text{OC}_2\text{F}_4\text{CO}_2\text{H}$
27. $\text{C}_4\text{F}_9\text{-O-(CF}_2\text{)}_3\text{CO}_2\text{H}$
28. $\text{C}_5\text{F}_{11}\text{-O-(CF}_2\text{)}_3\text{CO}_2\text{H}$
29. $\text{C}_6\text{F}_{13}\text{-O-(CF}_2\text{)}_3\text{CO}_2\text{H}$
30. $\text{C}_5\text{F}_{11}\text{-O-(CF}_2\text{)}_4\text{CO}_2\text{H}$
- 15 31. $\text{CF}_3\text{-O-(CF}_2\text{)}_5\text{CO}_2\text{H}$
32. $\text{C}_4\text{F}_9\text{-O-(CF}_2\text{)}_5\text{CO}_2\text{H}$
33. $\text{C}_5\text{F}_{11}\text{-O-(CF}_2\text{)}_5\text{CO}_2\text{H}$
34. $\text{C}_4\text{F}_9\text{-O-C}_4\text{F}_8\text{-O(CF}_2\text{)}_3\text{CO}_2\text{H}$
35. $\text{C}_6\text{F}_{13}\text{-O-C}_4\text{F}_8\text{-O(CF}_2\text{)}_3\text{CO}_2\text{H}$
- 20 36. $\text{C}_4\text{F}_9\text{-O-C}_2\text{F}_4\text{O-C}_2\text{F}_4\text{O(CF}_2\text{)}_3\text{CO}_2\text{H}$
37. $\text{CF}_3\text{-O-(C}_2\text{F}_4\text{O)}_3\text{-(CF}_2\text{)}_3\text{CO}_2\text{H}$
38. $\text{C}_8\text{F}_{17}\text{OCF}_2\text{OC}_5\text{F}_{10}\text{CO}_2\text{H}$
39. $(\text{CF}_3)_3\text{COC}_2\text{F}_4\text{OCF}_2\text{OC}_2\text{F}_4\text{OCF}_2\text{CO}_2\text{H}$
40. $(\text{CF}_3)_2\text{CF}_2\text{CF}_2\text{O(CF}_2\text{)}_5\text{CO}_2\text{H}$
- 25 41. $\text{CF}_3(\text{CF}_2)_7\text{OC}_2\text{F}_4\text{OC}_2\text{F}_4\text{OCF}_2\text{CO}_2\text{H}$
42. $\text{CF}_3(\text{CF}_2)_{11}\text{OC}_2\text{F}_4\text{OC}_2\text{F}_4\text{OCF}_2\text{CO}_2\text{H}$

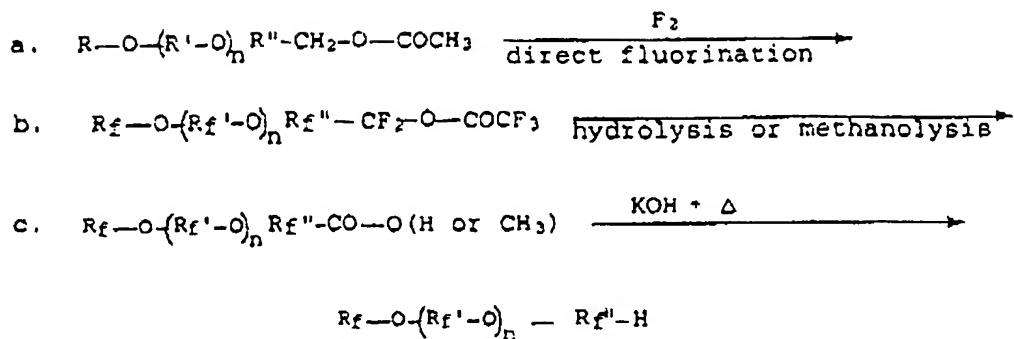
The following presents overall schemes of reactions that can be used in the preparation of omega-hydrofluoroalkyl ethers of this invention using general formulas defined above. In these schemes, the illustrated reaction results in the product whose formula is depicted on the succeeding line

Scheme I

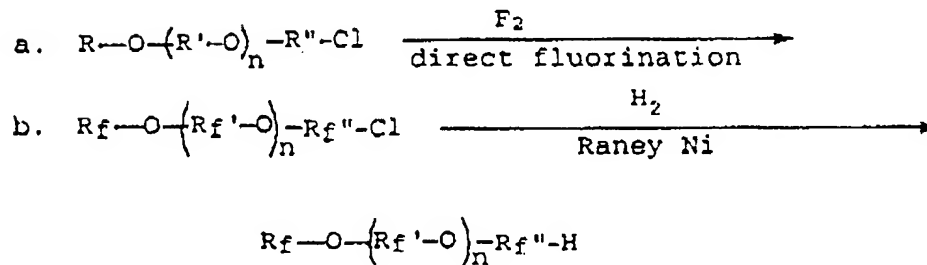


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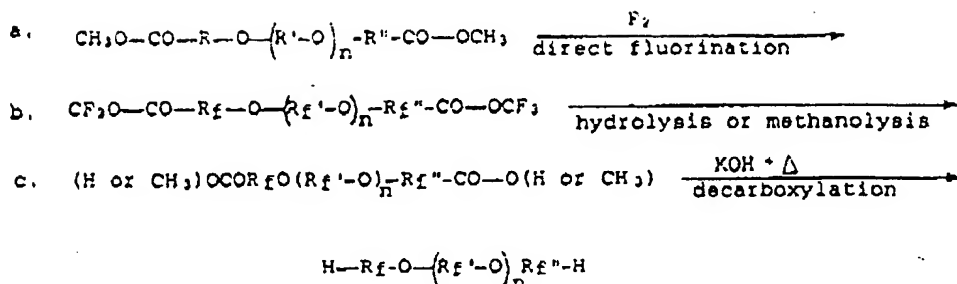
Scheme II

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Scheme III

10 The ether alpha and omega dihydrides, that is,
 where X in formula I is H, may be prepared by analogous
 schemes. For example, the following Scheme IV is
 analogous to Scheme I

15

Scheme IV

5 Looking first at Scheme I above, in the direct fluorination, step "a", a fluorinatable precursor ether carboxylic acid ester, e.g., $\text{C}_4\text{H}_9\text{-O-(CH}_2)_5\text{COOCH}_3$, is directly fluorinated by contact with fluorine gas. (The term "fluorinatable" means that the precursor contains carbon-bonded hydrogen atoms which are replaceable with fluorine and the precursor may contain unsaturation which can be saturated with fluorine.)

10 The resulting fluorinated ether acid ester compound, depicted in step b, can be made with essentially the same number and spatial arrangement of carbon and oxygen atoms as the precursor thereof. If a fluorinated ether acid composition which consists or consists essentially of a selected mixture of fluorinated ether compounds is desired, a selected

15 mixture of the corresponding precursor compounds can be fluorinated or, alternatively, the selected precursor compounds can be separately fluorinated and then blended.

20 The direct fluorination of the fluorinatable ether precursor can be carried out at temperatures typically used in direct fluorination, e.g., at moderate or near ambient temperatures such as -20°C to $+50^\circ\text{C}$, using a stoichiometric excess of fluorine gas, which is preferably diluted with an inert gas, such as nitrogen,

25

to minimize or avoid the hazards of pure fluorine gas and to control the amount of heat generated upon contact of the precursor with fluorine. The fluorination is preferably carried out in an oxygen- and water-free environment and can be carried out in the presence of solid, particulate scavenger, e.g., sodium fluoride, for the hydrogen fluoride by-product generated. Liquid phase direct fluorination can be employed and involves using an inert liquid, such as a fluorocarbon or chlorofluorocarbon liquid, as a reaction medium. Both scavenger and an inert liquid reaction medium can be utilized, if desired. The fluorination is preferably carried out by liquid phase direct fluorination in the absence of hydrogen fluoride scavenger by using a temperature and inert gas flow rate sufficient to volatilize hydrogen fluoride by-product and enable its removal from the fluorination zone as it is generated.

In another aspect, this invention provides a fluorochemical composition containing the fluorinated ether acid or derivative thereof, hereinbefore described, as the sole essential component of the fluorochemical composition.

Although direct fluorination is a substitution method involving the replacement of hydrogen atoms with fluorine, direct fluorination provides higher yields and purer products than do other substitution methods such as the electrochemical fluorination and cobalt trifluoride methods -- see, for example, U.S. Pat. No. 5,093,432 (Bierschenk et al.). The purity of the perfluorinated ether acid (or ester) compositions of the invention is further enhanced by the use of single precursor compounds or selected (rather than random) mixtures thereof.

The preferred method of fluorination is the "liquid phase direct fluorination technique," which involves making a very dilute dispersion or,

preferably, solution of the precursor(s) in a liquid reaction media, which is relatively inert to fluorine at the fluorination temperatures used, the concentration of fluorinatable starting material thus being relatively low so as to more easily control the reaction temperature. The reaction mixture can also contain or have dispersed therein a hydrogen fluoride scavenger such as sodium fluoride, the scavenger:precursor weight ratio being, for example, from about 0.5:1 to 7:1. The reaction mixture can be vigorously agitated while the fluorine gas is bubbled through it, the fluorine preferably being used in admixture with an inert gas, such as nitrogen, at a concentration of about 5 to 50 volume %, more preferably about 10 to 25 volume %, and being maintained in stoichiometric excess throughout the fluorination, e.g., up to 15 to 40%, or higher, depending on the particular starting material and the efficiency of the equipment used, such as the reactor agitation. Yields generally in the range of about 30-77 mole %, and, with experience, as high as 65 to about 85 mole %, of the perfluorinated product may be achieved by this method.

Suitable liquids useful as reaction media for the liquid phase direct fluorination technique are chlorofluorocarbons such as Freon™ 11 fluorotrichloromethane; chlorofluoroethers; Fluorinert™ electronic liquids FC-75, FC-72, and FC-40; perfluoroalkanes such as perfluoropentane and perfluorodecalin; perfluoropolyethers; and perfluoroacetals. Mixtures of such liquids can be used, e.g., to get good dispersion of precursor and intermediate reaction products. The reaction media are conveniently used at atmospheric pressure. Lower molecular weight members of the above classes of reaction media can also be used, but elevated pressures are then required to provide a liquid phase.

The liquid phase direct fluorination reaction is generally carried out at a temperature between about -10°C to +50°C, preferably between about -10°C to 0°C if a hydrogen fluoride scavenger is used, and, if such a scavenger is not used, between about 0°C to 150°C, preferably about 0°C to 50°C, most preferably about 10°C to 30°C, the temperature being sufficient to volatilize the hydrogen fluoride by-product and, with the aid of the inert gas, flowing at a sufficient rate, cause the purging of the by-product from the fluorination reactor as it is generated. At these temperatures, the liquids utilized as reaction media do not react appreciably with the diluted fluorine and are essentially inert. The reaction medium and other organic substances may to some extent be present in the gaseous reactor effluent, and a condenser may be used to condense the gaseous reaction medium and such substances in the effluent and permit the condensate to return to the reactor. The condenser can be operated so as to minimize or prevent the return to the reactor of hydrogen fluoride by-product (which could have an adverse effect on yield of product if allowed to remain in the reactor during fluorination). The return of the hydrogen fluoride can be minimized or prevented by selective condensation of the organic materials while allowing the hydrogen fluoride to pass through the condenser, or by total condensation of both hydrogen fluoride and the organic materials into a separate vessel and followed, if desired, by separation of the hydrogen fluoride as the upper liquid phase and the return of the lower liquid phase.

The liquid phase fluorination reaction may be carried out in a batch mode, in which all of the precursor is added to the liquid prior to fluorination to provide a precursor concentration of up to about 10% by weight, and the fluorine-containing gas is then bubbled through the precursor-containing liquid. The

reaction can also be carried out in a semi-continuous mode, in which the precursor is continuously pumped or otherwise fed neat, or as a diluted solution or dispersion, in a suitable liquid of the type used as a reaction medium, into the reactor, e.g., at a rate of about 1 to 3 g/hr into 400 mL of liquid reaction mixture, as fluorine is bubbled through, e.g., at a fluorine flow rate of about 40 to 120 mL/min and an inert gas flow rate of about 150 to 600 mL/min. The fluorination can also be carried out in a continuous manner, in which the precursor (either neat or dissolved or dispersed in a suitable liquid of the type used as a reaction medium) is continuously pumped or otherwise fed into the reactor containing the reaction medium as the fluorine-containing gas is introduced, as described above, and the stream of unreacted fluorine, hydrogen fluoride gas, and inert carrier gas is continuously removed from the reactor, as is a stream of liquid comprising perfluorinated product, incompletely fluorinated precursor, and inert liquid reaction medium, and the necessary separations are made to recover the fluoroalkyl ether composition. If desired, the unreacted fluorine and the incompletely fluorinated precursor can be recycled. The amount of inert liquid medium in the reactor can be maintained at a constant level by addition of recycled or fresh liquid.

Due to the extremely high exothermicity of the fluorination reaction, a cooled liquid or ice bath is generally employed in order that acceptable rates of reaction may be achieved. When the reaction is complete, the reactor is purged of fluorine and the reactor contents are removed. Where the fluorination is carried out by the liquid phase fluorination technique in the presence of a hydrogen fluoride scavenger, the spent scavenger can be separated by filtration or decantation from the liquid reactor

contents and the latter then distilled to separate the reaction medium from the crude product. Where the fluorination is carried out by the liquid phase fluorination technique without using the scavenger, the reaction product mixture can be distilled to recover the product.

Useful representative precursor fluorinatable ether acid esters which can be used to prepare the omega-hydrofluoroalkyl ethers of this invention are the hydrocarbon counterparts of the structures listed in Table A above, except that instead of the terminal hydrogen atom the structures of the esters terminate with $-Z'$ (where Z' is as defined for formula IV) or $-CH_2OC(O)R$ (as shown in Scheme II supra) and that the precursors can contain unsaturation.

Representative examples of the fluoroether acids of or used in this invention include the perfluorinated (i.e., having essentially all hydrogens replaced with fluorine) counterparts of the precursor fluorinatable acid esters described above. When the precursors have unsaturation, the corresponding fluorinated ether acids are saturated.

As pointed out above, the fluoroether acids and derivatives can be used as precursors in the preparation of the omega-hydrofluoroalkyl ethers and they are also useful, for example, as surfactants.

The above-described fluoroether acids or the esters thereof, e.g., alkyl esters such as the methyl ester, can be converted by a decarboxylation process to the omega-hydrofluoroalkyl ethers of this invention. In one such process, a solution of KOH in ethylene glycol is prepared and the fluoroether acid or ester precursor is added thereto (neat or as a solution in an inert solvent liquid such as a perfluorinated liquid), preferably dropwise with stirring at ambient or room temperature. The resulting mixture can then be heated slowly, for example, to 190°C , during which time the

methanol (from the saponification of a methyl ester), water (from neutralization of an acid), and decarboxylated product are distilled. The omega-hydrofluoroalkyl ethers of the invention are surprisingly stable under such harsh basic conditions. An inert solvent liquid, if used, can be removed, for example, at low temperature under vacuum after neutralization. The resulting distillate, comprising the omega-hydrofluoroalkyl ether product, can be washed with water, dried with silica gel or magnesium sulfate, and then distilled to purify the product. If desired, the hydrofluoroalkyl ether product can be refluxed with a solution of potassium permanganate in acetone to remove easily-oxidized impurities. The yields of the ether product are generally high and the product generally will be quite pure and consist or consist essentially of the desired omega-hydrofluoroalkyl ether.

The omega-hydrofluoroalkyl ether compositions are non-toxic and capable of dissolving and transporting oxygen and are therefore potentially useful as blood substitutes which can be employed invasively in the treatment of trauma, vascular obstructions, as adjuvants to cancer radiation treatment or chemotherapy, and as imaging contrast agents. For such uses, emulsions of the compositions can be prepared by methods such as those described, for example, in U.S. Pat. Nos. 3,911,138 (Clark) and 5,077,036 (Long) and the PCT International Application published as WO 93/11868 (Kaufman et al.). The omega-hydrofluoroalkyl ether compositions are also useful as solvents for cleaning and drying applications such as those described in U.S. Patent Nos. 5,125,089 (Flynn et al.), 3,903,012 (Brandreth), and 4,169,807 (Zuber). Minor amounts of optional components, e.g., surfactants, may be added to the fluoroether compositions to impart particular desired properties for particular uses. The

ether compositions are also useful as heat transfer agents or coolants in refrigerator or freezer compressors or air conditioners, blowing agents or cell size regulators in making polyurethane foam insulation, or chemical fire extinguishing agents in streaming applications, total flooding, explosion suppression and inertion, and as carrier solvents for highly fluorinated polyethers used as lubricants for magnetic recording media.

10 In using the omega-hydrofluoroalkyl ether compositions of this invention for the drying of or displacing water from the surface of articles, such as circuit boards, the processes of drying or water displacement described in U.S. Patent No. 5,125,978
15 (Flynn et al.) can be used. Broadly, such process comprises contacting the surface of an article with a liquid composition comprising the ether composition of this invention, preferably in admixture with a non-ionic fluoroaliphatic surface active agent. The wet
20 article is immersed in the liquid composition and agitated therein, the displaced water is separated from the liquid composition, and the resulting water-free article is removed from the liquid composition. Further description of the process and the articles
25 which can be treated are found in said U.S. Patent No. 5,125,978.

In using the ether composition of this invention as a heat transfer liquid in vapor phase soldering, the process described in U.S. Patent No. 5,104,034 (Hansen)
30 can be used. Briefly, such process comprises immersing the component to be soldered in a body of vapor comprising the ether composition of this invention to melt the solder. In carrying out such a process, a liquid pool of the ether composition of this invention
35 can be heated to boiling in a tank to form a saturated vapor in the space between the boiling liquid and a condensing means, a workpiece to be soldered is

immersed in the vapor whereby the vapor is condensed on the surface of the workpiece so as to melt and reflow the solder, and the soldered workpiece is then removed from the space containing the vapor.

5 In using the ether composition of this invention as a blowing agent in making plastic foam, such as foamed polyurethane, the process reactants, and reaction conditions described in U.S. Patent No. 5,210,106 (Dams et al.) can be used. In carrying out
10 such process, organic polyisocyanate and high molecular weight compound with at least 2 reactive hydrogen atoms, such as a polyol, are admixed in the presence of a blowing agent mixture comprising an ether composition of this invention, catalyst, and a surfactant.

15 This invention is further illustrated by the following examples, but the particular materials and amounts thereof recited in these examples, as well as other conditions and details, should not be construed to unduly limit this invention.

20 Example 1. Preparation of $C_8F_{17}-O-C_2F_4H$ from $C_8F_{17}-O-C_2F_4CO_2CH_3$

 The organic starting material, $C_8H_{17}-O-C_2H_4CO_2CH_3$,
25 was prepared by base-catalyzed Michael addition of n-octanol to acrylonitrile, followed by acid-catalyzed methanolysis. The methyl ester was directly fluorinated with F_2 to produce the fluorinated ester, $C_8F_{17}-O-C_2F_2CO_2CF_3$. This fluorination was carried out
30 in a 2-liter, jacketed reactor vessel of Monel™ metal equipped with a magnetic drive agitator, gas feed line, organic reactant feed line, and a reflux condenser. The gas feed line was 0.3 cm diameter tube reaching to a point below the bottom impeller of the agitator. The
35 feed line was a 0.15 cm diameter tube connected to a syringe pump. The reflux condenser consisted of about 6-meters of two coiled concentric tubes, the inner tube having a 1.27 cm diameter and the outer tube having a

2.54 cm diameter. Gases from the reactor were cooled in the inner tube by refrigerant, ethylene glycol-water, flowing in the annulus between the two tubes. The reactor was charged with about 1.8 liters of Freon™ 113 chlorofluorocarbon and purged with 650 mL/min of nitrogen for 20 minutes. The gas stream was then changed to a mixture of 310 mL/min fluorine and 1100 mL/min nitrogen. After about 12 minutes, 100 g of $C_8H_{17}-O-C_2H_4-CO_2CH_3$, diluted to 260 mL with Freon™ 113 chlorofluorocarbon, was fed to the reactor at a rate of 13 mL/hr (5 g/hr feed rate). The reactor contents were maintained at about 16-18°C throughout the fluorination. The condenser temperature was about -22°C. The fluorine flow was continued for ten minutes after complete addition of the organic feed. The reactor was then purged with nitrogen for one hour. The Freon™ 113 solution of the crude perfluorinated ester was treated with 150mL of 14% BF_3 in methanol and agitated vigorously for 24 hrs. The mixture was washed with water, dried over $MgSO_4$ and distilled (b.p. 40°C/0.2 torr) to yield $C_8F_{17}-O-C_2F_4-CO_2CH_3$ (47% yield). For purposes of decarboxylation, 39 g of 85% KOH was dissolved in approximately 300 mL of ethylene glycol and the above-described fluorinated methyl ester was added dropwise with stirring to the KOH solution at room temperature. Upon complete addition, the reaction mixture had a pH of 8 to 9. The mixture was heated slowly with stirring and the product of decarboxylation, $C_8F_{17}-O-C_2F_4H$, was distilled along with methanol from saponification of the methyl ester, water from the KOH and a small amount of ethylene glycol. When the reaction mixture temperature reached 170°C, the heating was stopped. The lower fluorochemical phase of the distillate was separated, washed with water, dried and distilled through a three-plate Snyder column. The main fraction, boiling at 146-150°C, yielded 122 g of product. Gas chromatography and mass

spectrometry (GC/MS) of a sample of the product showed the sample to be 94% pure and confirmed the structure as $C_8F_{17}-O-C_2F_2H$.

5 Example 2. Preparation of $C_8F_{17}-O-C_2F_4H$ from $C_8F_{17}-O-C_2F_4CO_2H$

10 $C_8H_{17}-O-C_2H_4CO_2CH_3$ was prepared by base-catalyzed Michael addition of n-octanol to acrylonitrile, followed by acid-catalyzed methanolysis. This carboxylic acid ester was directly fluorinated by essentially the same fluorination procedure described in Example 1 to produce the corresponding ether acid, $C_8F_{17}-O-C_2F_4COOH$ upon hydrolysis. Differential scanning calorimetry revealed multiple transitions, which is characteristic of polymorphism.

15 A solution of 116 g of 85% KOH in 800 mL of ethylene glycol was prepared in a 3 L round-bottom flask. 1000 g of the $C_8F_{17}OC_2F_4-CO_2H$ was added dropwise to the stirred KOH solution. Upon complete addition, an additional 10 g of KOH was added and the mixture heated. The fluorochemical product of decarboxylation was distilled together with a small amount of water from the neutralization of the acid. The lower fluorochemical phase of the distillate was separated, washed with salt water, dried over Na_2SO_4 and distilled as in Example 1 to yield 817 g of $C_8F_{17}-O-C_2F_4H$.

25 Example 3. Preparation of $C_7F_{15}-O-C_2F_4H$ from $C_7F_{15}-O-C_2F_4CO_2CH_3$

30 $C_7H_{15}-O-C_2H_4CO_2CH_3$ was prepared by base-catalyzed Michael addition of n-heptanol to acrylonitrile, followed by acid-catalyzed methanolysis. 550 g of the corresponding methyl ester, $C_7F_{15}-O-C_2F_4COOCH_3$, (prepared by essentially the same fluorination and methanolysis procedures of Example 1), was added dropwise to a solution of 166.6 g of KOH in approximately 880 mL of ethylene glycol. The

fluorochemical product was recovered essentially as in Example 1 to yield 440 g which was distilled through a six-plate Snyder column and the fraction boiling from 130 to 131°C was collected (340 g). This fraction was combined with 8.5 g of KMnO_4 and approximately 350 g of acetone and heated to reflux. After four hours, an additional 5 g of KMnO_4 was added and the resulting mixture was heated for an additional 3 hours. The mixture was filtered, the filter cake washed with acetone, and water was added to the filtrate causing a lower fluorochemical phase to form which was then washed with water, followed by conc. H_2SO_4 , again with water, and then filtered through silica. ^1H NMR and ^{19}F NMR confirmed the reaction product to have the desired structure, $\text{C}_7\text{F}_{15}\text{-O-C}_2\text{F}_2\text{H}$. Gas-liquid chromatography of a sample showed it to be 98.7% pure.

Example 4. Preparation of $\text{C}_6\text{F}_{13}\text{-O-C}_2\text{F}_4\text{-O-CF}_2\text{H}$ from $\text{C}_6\text{F}_{13}\text{-O-C}_2\text{F}_4\text{-OCF}_2\text{CO}_2\text{CH}_3$

The starting material, $\text{C}_6\text{H}_{13}\text{-O-C}_2\text{H}_4\text{-O-C}_2\text{H}_4\text{-O-COCH}_3$, was prepared by acetylation of hexyloxyethoxy ethanol with acetyl chloride. The acetate was then converted to $\text{C}_6\text{F}_{13}\text{-O-C}_2\text{F}_4\text{-OCF}_2\text{CO}_2\text{CH}_3$ by essentially the same fluorination and methanolysis procedures of Example 1. 548 g of this fluorochemical was combined with 144.2 g of KOH in 600 g of ethylene glycol. The resulting mixture was heated, distilled and the product, $\text{C}_6\text{F}_{13}\text{-O-C}_2\text{F}_4\text{-OCF}_2\text{H}$, was recovered as in Example 1. Total yield was 433 g. The product was again distilled (b.p 131°C) through a 12-inch (30.5 cm) perforated-plate column at atmospheric pressure. The structure of the product was confirmed by ^1H and ^{19}F NMR as $\text{C}_6\text{F}_{13}\text{-O-C}_2\text{F}_4\text{-OCF}_2\text{H}$. GC/MS revealed a sample of it to be 99.6% pure.

35

Example 5. Preparation of $\text{C}_8\text{F}_{17}\text{-O-CF}_2\text{H}$ from $\text{C}_8\text{F}_{17}\text{-O-CF}_2\text{-CO}_2\text{CH}_3$

$C_8H_{17}-O-C_2H_4-O-(CO)CF_3$ was prepared by acetylation of octyloxyethanol with trifluoroacetic anhydride. 100 g of the trifluoroacetate was directly fluorinated by essentially the same fluorination procedures of Example 1 and the fluorination product was quenched with a solution of BF_3 in methanol to yield crude $C_8F_{17}-O-CF_2-CO_2CH_3$, which was further purified by distillation, b. 92-97°C @20 torr.

A 58 g sample of the latter methyl ester was decarboxylated using 10.8 grams of KOH in ethylene glycol and the product, $C_8F_{17}-O-CF_2H$, was recovered as in Example 1. The structure of the product was confirmed by ^{19}F NMR. GLC revealed the product to be 99.6% pure, b. 134-136°C.

Example 6. Preparation of $C_4F_9-O-C_2F_4H$ from $C_4F_9-O-C_2F_4-CO_2CH_3$

The methyl ester, $C_4H_9-O-C_2H_4-CO_2CH_3$, was prepared by base-catalyzed Michael addition of n-butanol to acrylonitrile, followed by acid-catalyzed methanolysis. The methyl ester was then converted to the corresponding fluorinated methyl ester, $C_4F_9-O-CF_2CF_2-CO_2CH_3$, by essentially the same fluorination and methanolysis procedures described in Example 1.

1160 g of the latter methyl ester was added dropwise with stirring to 3103 g of ethylene glycol and 129.5 g of NaOH. The product was distilled (b.p. 83°C) and treated with $KMnO_4$ /acetone, and worked up as in Example 3. The structure of the purified compound, $C_4F_9-O-CF_2CF_2H$, was confirmed by 1H and ^{19}F NMR and GC/MS.

A sample of this compound was evaluated for use in precision cleaning applications by measuring the solubilities of selected hydrocarbon solvents in the sample. High solubility would indicate improved performance as a cleaning agent relative to perfluorocarbon solvents. The following hydrocarbon

solvents were found to be soluble in amounts up to 50% by weight with the ether hydride: hexane, heptane, toluene, acetone, 2-butanone, 4-methyl-2-pentanone, ethyl acetate, methanol, ethanol, isopropanol, dimethyl
5 formamide, trans-1,2-dichloroethylene and isopropyl ether. o-Xylene was found to be soluble to 19% by weight. Chloroform was found to be soluble to 45% by weight. Ethylene glycol was found to be soluble to less than 15% by weight and a light hydrocarbon oil was
10 found to be soluble to less than 0.05% by weight.

A sample of the compound was also evaluated for use in spot-free drying applications such as taught in U.S. Patent No. 5,125,978 (Flynn et al.). A water displacement composition was prepared by dissolving
15 0.2% by weight of $C_4F_9OC_2F_4OCF_2CONHC_2H_4OH$ in $C_4F_9-O-C_2F_4H$. The solution was heated to 45°C in an ultrasonic bath. Using the procedure described in U.S. Patent No. 5,125,978, test coupons of glass and stainless steel were wetted with water and subsequently immersed into
20 this solution with ultrasonic agitation. All water was displaced within 60 seconds.

A sample of this compound was also evaluated for use as a rinse agent in co-solvent cleaning applications. (Such cleaning applications are taught,
25 for example, in International Patent Publication No. WO 92/22678 (Petroferm Inc.). Organic esters such as methyl decanoate have found utility as solvating agents in two-phase cleaning applications using perfluorohexane as the carrier liquid and rinse agent.)
30 Methyl decanoate and $C_4F_9OC_2F_4H$ were placed in separate containers and heated to 50°C in an ultrasonic bath. A 50 mm X 25 mm X 1.5 mm aluminum coupon was contaminated with 0.0831 g of a light hydrocarbon oil. The contaminated coupon was first immersed in the methyl
35 decanoate for about 60 seconds and then immersed in the $C_4F_9OC_2F_4H$ for about 60 seconds. The $C_4F_9OC_2F_4H$ rinsed 100 percent (as determined by weight difference) of the

oil and the methyl decanoate from the coupon. Under the same conditions, perfluorohexane removed only 98.5 percent of the oil and methyl decanoate, indicating that $C_4F_9OC_2F_4H$ can be more effective as a carrier liquid and rinse agent than perfluorohexane.

Example 7. Preparation of $HCF_2CF_2-O-CF_2CF_2-O-CF_2CF_2H$ from $CH_3OC(O)C_2F_4-O-C_2F_4-O-C_2F_4C(O)OCH_3$

10 The starting material, $CH_3OC(O)C_2H_4-O-C_2H_4-O-C_2H_4C(O)OCH_3$, was prepared by base-catalyzed Michael addition of ethylene glycol to acrylonitrile, followed by acid-catalyzed methanolysis. The starting material was then fluorinated and methanolysed by essentially
15 the same procedures described in Example 1 to give $CH_3OC(O)C_2F_4-O-C_2F_4-O-C_2F_4C(O)OCH_3$.

1136 grams of the fluorinated ester was added to a mixture of 305.6 g of KOH in 2665 g of ethylene glycol. The decarboxylation was carried out essentially as
20 described in Example 1, and the crude product distilled after phase separation but without water washing. The distillate still contained methanol which was removed by a wash with concentrated sulfuric acid followed by two water washes to give 695 g of the desired ether
25 hydride product, with a boiling range of 93-94°C.

Example 8. Preparation of $C_4F_9-O-(CF_2)_5H$ from $C_4F_9-O-(CF_2)_5-CO_2H$

30 118.2 g (1.0 mol) hexane-1,6-diol, 4.4 g Adogen™ 464 quaternary ammonium salt, 80.0 g (2.0 mol) NaOH, and 250 mL tetrahydrofuran was stirred at reflux. 80 mL H_2O was added to get better mixing. After 20 min more, 137 g (1.0 mol) butyl bromide was added over 0.5
35 hr, and stirred overnight at reflux. The reaction mixture was quenched in 1 L H_2O , and the upper layer was combined with an ether extract of the lower layer, dried over $MgSO_4$, and stripped on a rotary evaporator. Treating the resulting stripped layer (151 g) in 100 mL
40 $CHCl_3$ with 150 mL acetyl chloride added dropwise and

subsequently heating at reflux 4 hr and solvent removal gave 225.4 g of liquid. Distillation of the liquid gave 176.0 g (b. 100-104°C/0.9 torr) of distillate. GLC indicated 56% of it to be the desired 6-butoxyhexyl acetate, accompanied by hexanediol diacetate and dibutoxyhexane. 100 g of this mixture was fluorinated essentially as in Example 1. Treatment of the resulting fluorinated product with 30 mL of a 10 weight percent solution of H_2SO_4 in H_2O and shaking at room temperature for 2 hours, filtration of solid fluorinated adipic acid, separation of the F-113 layer, drying over MgSO_4 , and distillation produced a main cut of 73.4 g, b. 116°C/20 torr, 96% pure $\text{C}_4\text{F}_9\text{-O-(CF}_2)_5\text{COOH}$. The latter was added to a solution of 10.0 g (0.25 mol) NaOH and 100 mL ethylene glycol and the mixture was heated to 120°C, with $\text{C}_4\text{F}_9\text{-O(CF}_2)_6\text{-O-C}_4\text{F}_9$ impurity from fluorination collecting in the Dean-Stark trap. On continued heating, gas evolution began and a liquid, $\text{C}_4\text{F}_9\text{-O(CF}_2)_5\text{H}$, (44.6 g) collected in the trap, ending by 170°C. The collected liquid was dried over silica gel and distilled on a 4-inch (10.2 cm) Vigreux column to 38.8 g, b.p 131°C. F-nmr confirmed structure, in high purity, to be $\text{C}_4\text{F}_9\text{-O-(CF}_2)_5\text{H}$.

Example 9. Preparation of $\text{C}_5\text{F}_{11}\text{-O-(CF}_2)_5\text{H}$ from $\text{C}_5\text{F}_{11}\text{-O-(CF}_2)_5\text{COOH}$

In a similar fashion to Example 8, hexanediol was alkylated with n-pentyl bromide, the product was acetylated, and the crude acetate, $\text{C}_5\text{H}_{11}\text{-O-(CH}_2)_6\text{OC(O)CH}_3$, was distilled (b. 125°C/3 torr) and the distillate was fluorinated essentially by the fluorination procedure of Example 1. The fluorinated ester was hydrolyzed to the corresponding acid. Decarboxylation of the fluorinated acid, $\text{C}_5\text{F}_{11}\text{O(CF}_2)_5\text{COOH}$, with NaOH gave 829 g of product. The product was washed with water, dried over MgSO_4 , and

distilled to yield 555 g of $C_5F_{11}-O-(CF_2)_5H$, b. 145-149°C.

5 Example 10. Preparation of $C_8F_{17}-O-(CF_2)_5H$ from $C_8F_{17}-O-(CF_2)_5COOH$

10 In a fashion similar to Example 8, hexanediol was alkylated with n-octyl bromide, the product was acetylated, and the resulting $C_8H_{17}-O-(CH_2)_6-O-COCH_3$ was directly fluorinated and hydrolyzed as in Example 8 to $C_8F_{17}-O-(CF_2)_5COOH$, which was recrystallized from perfluorohexane. The recrystallized acid (37.5 g) was mixed with 4.0 g NaOH and 100 mL ethylene glycol and heated to 185°C. The product was washed with water, and the residual 27.9 g was distilled to give pure $C_8F_{17}-O-(CF_2)_5H$, micro b.p. 195°C.

20 Example 11. Preparation of $C_4F_9-O-CF_2C(CF_3)_2CF_2H$ from $C_4F_9-O-CF_2C(CF_3)_2CF_2Cl$

25 The alkylation of 2,2-dimethyl-1,3-propanediol with n-butyl bromide carried out essentially as in Example 8 gave the crude mono-alkylated product, which was treated with $SOCl_2$ to give $C_4H_9-O-CH_2C(CH_3)_2CH_2Cl$, b. 80-90°C/20-30 torr. This compound was then fluorinated as in Example 1 to give $C_4F_9-O-CF_2C(CF_3)_2CF_2Cl$. 20.0 g of the latter chloride was mixed with 5.3 g water-wet Raney Ni and 50 mL of NH_3 -saturated methanol. The mixture was left shaking on a Parr hydrogenation apparatus for 3 days at about 25 °C, with most of the 21 kPa (3 psig) hydrogen pressure drop occurring in the first day. The product was recovered by filtration and quenched in water, yielding 7.9 g with some mechanical loss. ^{19}F -nmr confirmed the product to be $C_4F_9-O-CF_2C(CF_3)_2CF_2H$. A scaleup to 100 g gave 47 g, distilled to b.p 135°C.

Example 12. Preparation of $\text{H}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{H}$ from $\text{Cl}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{Cl}$

5 $\text{Cl-}(\text{CH}_2)_4\text{-O-}(\text{CH}_2)_4\text{-Cl}$ was fluorinated as in Example 1 to provide $\text{Cl}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{Cl}$. A mixture of 30.3 g $\text{Cl}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{Cl}$, 11.3 g fresh water-wet Raney Ni, and 200 mL methanol was purged for several minutes with NH_3 and pressurized with 310 kPa (45 psig) hydrogen on a Parr hydrogenation apparatus at about 25 °C. After 17
10 hr, pressure had dropped to 255 kPa (37 psig) and the mixture had become acidic, with glass etching noted. More ammonia was added and the reduction was continued, dropping another 62 kPa (9 psig). The reaction product was filtered and quenched in water to give 15.4 g of
15 lower phase, 68% pure product confirmed by GLC to be $\text{H}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{H}$. Distillation yielded 27.0 g, b. 121-124 °C, 87% pure.

20 Example 13. Preparation of $\text{H}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{H}$ and $\text{Cl}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{H}$ from $\text{Cl}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{Cl}$

A mixture of 50.0 g $\text{Cl}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{Cl}$ and 30 g Zn in butanol was stirred at 110 °C for 2 days. GLC of a sample of the resulting reaction product indicated
25 partial conversion. 21 g more Zn was added and the mixture was heated one more day. Filtration and quenching of the resulting material in water gave 27.0 g of a colorless liquid. The product was 35% of
30 $\text{H}(\text{CF}_2)_4\text{-O-}(\text{CF}_2)_4\text{H}$, 42% mono hydride, and 16% unreduced dichloride.

Example 14. Preparation of $\text{C}_6\text{F}_{13}\text{-O-CF}_2\text{CF}_2\text{H}$ from $\text{C}_6\text{F}_{13}\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$

35 The starting material, $\text{C}_6\text{H}_{13}\text{-O-C}_2\text{H}_4\text{-CO}_2\text{CH}_3$, was prepared by the Michael addition of hexanol to acrylonitrile followed by acid-catalyzed esterification with methanol. The resulting ester was then fluorinated and hydrolyzed to give the $\text{C}_6\text{F}_{13}\text{-O-C}_2\text{F}_4\text{CO}_2\text{H}$.

500 g of the acid $C_6F_{13}-O-C_2F_4CO_2H$, was added slowly to a solution of 68.7 g KOH in 700 g ethylene glycol. At the end of the addition, an additional 5 g of KOH was added to the homogeneous solution to bring the pH to 9. The decarboxylation was carried out as in Example 1 and subsequently distilled, producing 327 g of product, b. 104-107 °C. The product was treated with potassium permanganate essentially as in Example 3. GC/MS, ^{19}F nmr, 1H nmr and IR confirmed structure of the product as $C_6F_{13}-O-CF_2CF_2H$.

Example 15. Preparation of $C_6F_{13}-O-CF_2H$ from $C_6F_{13}-O-CF_2CO_2CH_3$

The starting material, $C_6H_{13}-O-C_2H_4OC(O)CH_3$, prepared by acetylation of ethylene glycol monohexyl ether, was fluorinated and decarboxylated by essentially the procedures of Example 1 to give 146 g of $C_6F_{13}-O-CF_2H$ (b. 92-96°C).

Example 16. Preparation of $CF_3CF(CF_3)CF_2-O-CF_2H$ from $CF_3CF(CF_3)CF_2-O-CF_2CO_2CH_3$

The starting material, $CH_3CH(CH_3)CH_2-O-CH_2CH_2-OC(O)CH_3$, was prepared by acetylation of ethylene glycol monoisobutyl ether and conversion by essentially the fluorination and methanolysis procedures of Example 1 to give the methyl ester, $CF_3CF(CF_3)CF_2-O-CF_2CO_2CH_3$, b. 118-120°C.

149 g of the methyl ester was added to 28.6 g of KOH in 700 g of ethylene glycol rapidly dropwise. The decarboxylation was carried out to afford, after distillation, the product cut, 70 g, b. 45-47°C, of 99% purity by GLC. The structure was confirmed by GC/MS, 1H nmr, and ^{19}F nmr analysis as $CF_3-CF(CF_3)CF_2-O-CF_2H$.

Example 17. Preparation of $C_4F_9-O-(CF_2)_4-O-(CF_2)_3H$ from $C_4F_9-O-(CF_2)_4-O-(CF_2)_3COOCH_3$

The starting material, $C_4H_9-O-C_4H_8-O-(CH_2)_3CH_2OCOCH_3$, was directly fluorinated and methanolysed essentially by the procedures of Example 1 to produce $C_4F_9-O-C_4F_8-O-(CF_2)_3CO_2CH_3$. 56 g of the latter was added rapidly to a solution of 5.6 g KOH in 250 ml of ethylene glycol. The decarboxylation was carried out and the product phase separated, washed once with brine, and distilled to yield 36.6 g of product (b.p. 155-158°C) of GLC purity 100%. GC/MS, 1H , and ^{19}F nmr analysis confirmed the product to be $C_4F_9-O-(CF_2)_4-O-(CF_2)_3H$.

Example 18. Preparation of $(C_2F_5)_2CFCF_2-O-C_2F_4H$ from $(C_2F_5)_2CFCF_2-O-CF_2CF_2-C(O)OCH_3$

Starting material, $(C_2H_5)_2CHCH_2-O-CH_2CH_2C(O)OCH_3$, prepared by the Michael addition of 2-ethylbutanol to acrylonitrile followed by acid-catalyzed esterification with methanol, was fluorinated and methanolysed essentially by the procedures of Example 1 to give $(C_2F_5)_2CFCF_2-O-CF_2CF_2-C(O)OCH_3$, b.p. 159°C, the direct fluorination yield, based on the methyl ester starting material being 88%.

The decarboxylation was carried out essentially as in Example 1 and the product distilled at 108-110°C to yield 145 g, the IR analysis of which was consistent with the structure $(C_2F_5)_2CFCF_2-O-CF_2CF_2H$.

Example 19. Preparation of $c-C_6F_{11}CF_2-O-C_2F_4H$ from $c-C_6F_{11}CF_2-O-C_2F_4C(O)OCH_3$

The starting material, $c-C_6H_{11}CH_2-O-C_2H_4C(O)OCH_3$, prepared by the reaction of cyclohexylmethanol with acrylonitrile followed by acid-catalyzed esterification with methanol, was then fluorinated and methanolysed with BF_3 in methanol by essentially the procedures of Example 1 to give a 65% yield (based on the fluorination) of $c-C_6F_{11}CF_2-O-C_2F_4C(O)OCH_3$.

224 g of the latter fluorinated ester was added to a solution of 28.2 g of 85% KOH and 466 g ethylene glycol held at 60°C. The resulting mixture was then heated to 100°C and its pH adjusted to a pH greater than 7 by the addition of 5 g of 45 wt% aqueous KOH. Decarboxylation was carried out by distillation of the resulting mixture. The lower fluorochemical phase of the resulting distillate was separated therefrom, washed with an equal volume of water, and distilled at 123-126°C to give 155 g of a product (99.7% purity). The product was treated with KMnO_4 in acetone to give $\text{C-C}_6\text{F}_{11}\text{CF}_2\text{-O-C}_2\text{F}_4\text{H}$.

Example 20. Preparation of $\text{C}_4\text{F}_9\text{-O-C}_2\text{F}_4\text{-O-C}_3\text{F}_6\text{H}$ from $\text{C}_4\text{F}_9\text{-O-C}_2\text{F}_4\text{-O-C}_3\text{F}_6\text{C(O)OCH}_3$

$\text{C}_4\text{H}_9\text{-O-C}_2\text{H}_4\text{-O-C}_4\text{H}_8\text{OC(O)CH}_3$ was fluorinated and methanolysed by essentially the procedure of Example 1. The resulting product, $\text{C}_4\text{F}_9\text{-O-C}_2\text{F}_4\text{-O-C}_3\text{F}_6\text{C(O)OCH}_3$, in the amount of 419 g was rapidly added dropwise to a mixture of 49.4 g KOH in 800 g ethylene glycol. The resulting mixture was then heated slowly to a final flask temperature of 190°C. During such heating, methanol from the saponification of the ester, water, and $\text{C}_4\text{F}_9\text{-O-C}_2\text{F}_4\text{-O-C}_3\text{F}_6\text{H}$ distilled from the reaction mixture. Water was added to the distillate and the lower, fluorochemical phase (355 g) was separated and distilled (b. 120-122°C) to provide 308 g $\text{C}_4\text{F}_9\text{-O-C}_2\text{F}_4\text{-OC}_3\text{F}_6\text{H}$ (82% yield).

Example 21. Preparation of $\text{C}_6\text{F}_{13}\text{-O-C}_4\text{F}_8\text{-H}$ from $\text{C}_6\text{F}_{13}\text{-O-C}_4\text{F}_8\text{-CO}_2\text{CH}_3$

The starting material, $\text{C}_6\text{H}_{13}\text{-O-C}_5\text{H}_{10}\text{-OC(O)CH}_3$, was prepared by monoalkylation of 1,5-pentanediol with hexyl bromide, followed by acetylation with acetyl chloride. This compound was fluorinated and methanolysed by essentially the procedure of Example 1,

to give $C_6F_{13}-O-C_4F_8-CO_2CH_3$, b.p. $100^\circ C$ @ 13 torr. This ester was decarboxylated by heating a solution of 200 grams of ester in 250 mL of ethylene glycol with 30 g of KOH until the hydride product distilled. This liquid was washed with water, dried over $MgSO_4$ to give 128 g of $C_6F_{13}-O-C_4F_8-H$ of 82% purity. This was further purified by distillation using a 12 plate packed glass column, b.p. $146^\circ C$. The structure was confirmed by ^{19}F NMR.

Example 22. Preparation of $C_6F_{13}-O-C_3F_6-H$ from $C_6F_{13}-O-C_3F_6-CO_2^-K^+$

The starting material, $C_6H_{13}-O-C_4H_8-OC(O)CH_3$, was prepared by monoalkylation of 1,4-butanediol with hexyl bromide, followed by acetylation with acetic anhydride. This compound was fluorinated and methanolysed by essentially the procedure of Example 1, to give $C_6F_{13}-O-C_3F_6-CO_2CH_3$. The methyl ester was saponified using excess KOH, and then dried in a vacuum oven to yield the potassium salt. 575 g of the salt was heated with stirring in 250 mL of ethylene glycol and the product hydride recovered from the distillate, b.p. $129^\circ C$. The structure was confirmed by ^{19}F NMR.

Example 23. Preparation of $C_5F_{11}-O-C_4F_8-H$ from $C_5F_{11}-O-C_4F_8-CO_2^-Na^+$

The starting material, $C_5H_{11}-O-C_5H_{10}-O-C(O)CH_3$ was prepared by monoalkylation of 1,5-pentanediol with pentyl bromide, followed by acetylation with acetyl chloride. This compound was fluorinated and methanolysed by essentially the procedure of Example 1, to give $C_5F_{11}-O-C_4F_8-CO_2CH_3$. The methyl ester was saponified using excess NaOH, and decarboxylated and distilled essentially as in Example 22. Distillation through a twelve-plate packed glass column gave pure $C_5F_{11}-O-C_4F_8-H$, b.p. $125^\circ C$. The structure was confirmed by ^{19}F NMR.

Example 24. Preparation of $C_4F_9-O-C_3F_6-H$ from $C_4F_9-O-C_3F_6-CO_2^-Na^+$

The starting material, $C_4H_9-O-C_4H_8-OC(O)CH_3$, was prepared by monoalkylation of 1,4-butanediol with butyl bromide, followed by acetylation with acetyl chloride. This compound was fluorinated and methanolysed by essentially the procedure of Example 1, to give $C_4F_9-O-C_3F_6-CO_2CH_3$. This methyl ester was saponified, decarboxylated and the crude hydride recovered as in Example 23, and then further distilled to yield pure $C_4F_9-O-C_3F_6-H$, b.p. $90^\circ C$. The structure was confirmed by ^{19}F NMR.

Example 25. Evaluation of surfactant activity of perfluoroether carboxylic acids.

The surfactant activity of novel perfluoroether carboxylic acids of this invention was measured with a DeNuoy tensiometer after conversion of the acids to the corresponding ammonium salts. The acids were prepared by direct fluorination of their hydrocarbon precursors, followed by hydrolysis. The ammonium salts were prepared by treatment of the acid with excess aqueous ammonia followed by freeze drying. The results are reported in dynes/cm in the following Table C which lists the parent acid (from Table B) of the ammonium salt.

TABLE C

Parent Acid from Table B	Melting Point(s) of Acid ($^\circ C$)	Surface Tension (dynes/cm)			
		Concentration of Ammonium Salt			
		50 ppm	100 ppm	500 ppm	1000 ppm
18	-1	24	21	18	
2		21	24		
4		63	59	39	29
5		33	33	26	22
7	19	37	26	19	17

Parent Acid from Table B	Melting Point(s) of Acid (°C)	Surface Tension (dynes/cm)			
		Concentration of Ammonium Salt			
		50 ppm	100 ppm	500 ppm	1000 ppm
6		31	23	18	17
8		33	31	26	24
9		38	24	20	18
10		38	35	24	19
27	-39	67	63	50	43
29	-9	49	43	29	23
35		18	18	16	16
25	-9	19	15	15	14
30		46	43	32	24
31		69	68	48	52
32		63	54	31	21
33		28	21	15	
40		48	41	23	19
11	49, 59	18	15	15	
14		39	31	18	
15	16, -27	30	17	17	
16		24	19	18	17
17	90	14	15	15	

Example 26. Evaluation of ethers as fire extinguishing agents.

Omega-hydrofluoroalkyl ether compounds of this invention were evaluated as fire extinguishing agents using the National Fire Protection Association 2001 Fire Protection Standard, with a cup burner modified to handle liquid compounds. The results, shown below in Table D, indicate that the compounds could be effective agents for fire extinguishing, explosion suppression, and as flammable atmosphere inerting agents.

TABLE D

Agent	Extinguishment concentration, vol. %
$C_4F_9OC_2F_4H$	5.6
$HC_3F_6OC_3F_6H$	5.7

5

Example 27. Preparation of foamed polyurethane.

Omega-hydrofluoralkyl ether compounds of this invention were evaluated as blowing agents for foams using the procedures taught in U.S. Patent No. 5,210,106 (Dams et al.). Component A contained 15.0 parts by weight of PAPI™27, a methylene diphenyldiisocyanate having an isocyanate equivalent of 134.0, available from Dow Chemical. Component B of the foam contained 10.5 parts by weight (pbw) of Voranol™ 360, a polyether polyol with a hydroxyl number of 360 available from Dow Chemical; 0.26 pbw of water; 0.26 pbw of an oligomeric fluorochemical surfactant as described in Example 1 of U.S. Patent No. 3,787,351; 0.13 pbw of Polycat™ 8, a N,N-dimethylcyclohexylamine catalyst available from Air Products; and 1.87 pbw of $C_4F_9OCF_2CF_2H$ as the blowing agent.

The ingredients of Component B were mixed to obtain an emulsion which was then admixed with Component A and stirred at 2500 rpm for 10 seconds. The cream time of the foam was approximately 10 seconds. Rise time and tack-free time was approximately 2 and 3 minutes respectively. The

resulting polyurethane foam was rigid and had a uniform distribution of very fine, closed cells.

Example 28. Preparation of $(\text{CF}_3)_3\text{COC}_2\text{F}_4\text{OCF}_2\text{OCF}_2\text{CO}_2\text{CH}_3$.

5 The precursor, $(t\text{-C}_4\text{H}_9\text{OC}_2\text{H}_4\text{O})_2\text{CH}_2$, prepared by alkylation of methylene chloride with t-butoxy ethanol, was fluorinated and methanolysed essentially as in Example 1 to yield $(\text{CF}_3)_3\text{COC}_2\text{F}_4\text{OCF}_2\text{OCF}_2\text{CO}_2\text{CH}_3$, having a boiling range 80-82°C at 18 torr, and whose structure
10 was confirmed by ^{19}F NMR.

Example 29. Preparation of $\text{C}_8\text{F}_{17}\text{OCF}_2\text{OC}_3\text{F}_6\text{H}$ from $\text{C}_8\text{F}_{17}\text{OCF}_2\text{OC}_3\text{F}_6\text{CO}_2\text{CH}_3$.

15 The precursor, $\text{C}_8\text{H}_{17}\text{OCH}_2\text{OC}_4\text{H}_8\text{OH}$ was prepared by monoalkylation of butane diol with octyl chloromethyl ether. The precursor was first acetylated with acetyl chloride in methylene chloride containing triethylamine and then fluorinated, and a portion of the crude perfluorinated product was hydrolyzed by treatment with
20 aqueous sulfuric acid and then distilled to yield the carboxylic acid $\text{C}_8\text{F}_{17}\text{OCF}_2\text{OC}_3\text{F}_6\text{CO}_2\text{H}$, having a boiling range 100-106°C at 1.1 torr. Differential scanning calorimetry revealed the acid had a T_g of
25 -97.0°C and several crystalline exotherms of -77.4, -61.5 and -37.7°C and a broad melting point at -9.0°C.

30 Another portion of the crude perfluorinated products was methanolysed essentially as in Example 1 to yield $\text{C}_8\text{F}_{17}\text{OCF}_2\text{OC}_3\text{F}_6\text{CO}_2\text{CH}_3$, having a boiling range 124-130°C at 25 torr. The latter methyl ester was then decarboxylated using the procedure of Example 1 to
35 yield $\text{C}_8\text{F}_{17}\text{OCF}_2\text{OC}_3\text{F}_6\text{H}$, having a boiling range of 178-183°C; the structures of this hydride and the precursor fluorinated ester were confirmed by ^{19}F NMR.

Example 30. Preparation of $\text{C}_8\text{F}_{17}\text{O}-(\text{C}_2\text{F}_4\text{O})_2\text{CF}_2\text{CO}_2\text{H}$.

The precursor was prepared by monoalkylation of triethylene glycol with octyl bromide, followed by

acetylation. The precursor was fluorinated as in Example 1, hydrolyzed by treatment with aqueous sulfuric acid, and distilled, the product, $C_8F_{17}O-(C_2F_4O)_2CF_2CO_2H$, having a boiling range of 105-110°C at 1.4 torr, and a melting point of 24°C.

Example 31. Preparation of $HC_3F_6OC_3F_6H$ from $CH_3O(CO)C_3F_6OC_3F_6COOCH_3$

The starting diacetate, $CH_3C(O)OC_4H_8O-(C_4H_8O)_n C_4H_8OC(O)CH_3$, was prepared by acetylation of polytetramethylene glycol (average molecular weight of 250) with acetyl chloride. The diacetate was then converted to $CH_3OC(O)C_3F_6O-(C_4F_8O)_n C_3F_6COOCH_3$ by essentially the same fluorination and methanolysis procedures described in Example 1. 1400 g of the resulting mixture of diesters was distilled on a ten-plate glass-packed column to isolate $CH_3OC(O)C_3F_6OC_3F_6COOCH_3$.

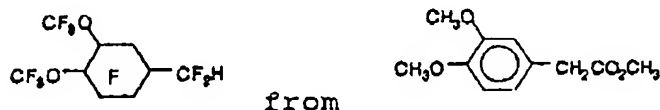
278 g of the isolated fluorochemical was combined with 72 g of KOH in 250 mL of ethylene glycol. The resulting mixture was heated, distilled, and the product, $HC_3F_6OC_3F_6H$, was recovered essentially as in Example 1 (b.p. 84°C). The structure of the product was confirmed by ^{19}F NMR.

Example 32. Preparation of $n-C_{12}F_{25}OC_2F_4OC_2F_4OCF_2CO_2H$.

The precursor, $n-C_{12}H_{25}O(C_2H_4O)_3H$, was prepared by monoalkylation of triethylene glycol with n-dodecyl bromide. After acetylation, the resulting product was fluorinated essentially as in Example 1, and the fluorinated product was concentrated and treated with 55.0 g NaOH in 300 mL water. After heating for 5 hours on a steam bath, the product was acidified with an excess of a 50 weight percent solution of H_2SO_4 in water and then extracted with FluorinertTM FC-75

perfluorinated liquid (a mixture of C₈ perfluorochemicals, b.p. 103°C) which had been heated to about 60°C on a steam bath. Distillation yielded pure n-C₁₂F₂₅OC₂F₄OC₂F₄OCF₂CO₂H (T_g = -62.7°C and T_m = 69.2°C by DSC).

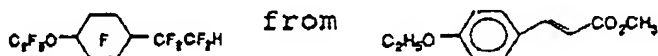
Example 33. Preparation of



10

The starting material, methyl 2-(3,4-dimethoxyphenyl)acetate was fluorinated essentially as in Example 1 to yield perfluoro-2-(3,4-dimethoxycyclohexyl)acetic acid after hydrolysis. This was then decarboxylated essentially as described in Example 1 to the perfluorinated ether hydride.

Example 34. Preparation of

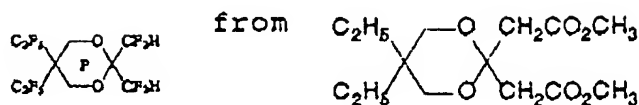


20

The starting material, methyl 3-(4-ethoxyphenyl)-trans-2-propenoate was prepared by condensation of 4-ethoxybenzaldehyde with malonic acid, followed by esterification. This methyl ester was fluorinated, methanolized, and decarboxylated essentially as in Example 1 to produce the perfluorinated ether hydride.

30

Example 35. Preparation of



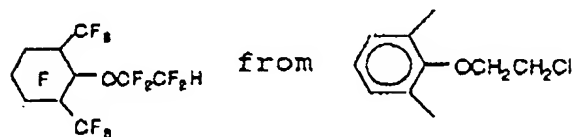
5 The starting material was prepared by condensation
of 2,2-diethyl propane diol with dimethyl 3-
oxoglutarate. This dimethyl ester was fluorinated,
methanolized to the diester, and decarboxylated
10 essentially as in Example 1 to produce the
perfluorinated ether dihydride.

Example 36. Preparation of



15 The starting material was prepared by reaction of
2,6-dimethylphenol with ethylene carbonate and
subsequent acetylation with acetyl chloride. This
acetate was fluorinated, methanolized, and
20 decarboxylated essentially as in Example 1 to produce
the perfluorinated ether hydride (b.p. 132°C).

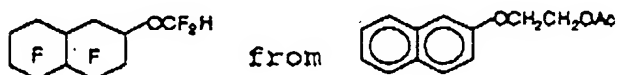
Example 37. Preparation of



25

The starting material was prepared by the treatment of 2-(2,6-dimethylphenyloxy)ethanol (from Example 36) with thionyl chloride. This was fluorinated essentially as in Example 1, followed by Raney Ni reduction of the chloride essentially as described in Example 12 to produce the perfluorinated ether hydride.

Example 38. Preparation of



The starting material was prepared from the addition of β -naphthol to ethylene carbonate, followed by acetylation with acetyl chloride. This acetate was fluorinated, methanolized, and decarboxylated essentially as in Example 1 to produce the perfluorinated ether hydride (b.p. 171°C).

Example 39. Preparation of $C_5F_{11}OCF_2C(CF_3)_2CF_2H$ from $C_5H_{11}OCH_2C(CH_3)_2CH_2Cl$

The starting material was prepared essentially as described above in Example 11. The ether chloride was fluorinated essentially as in Example 1, followed by Raney Ni reduction of the chloride essentially as described in Example 11 to produce the perfluorinated ether hydride (b.p. 148°C).

Example 40. Preparation of $(C_4F_9O)_2CFCF_2H$ from $(C_4H_9O)_2CHCH_2Cl$

The starting material was prepared by the addition of n-butanol to 2-chloroacetaldehyde and was fluorinated essentially as in Example 1, followed by Raney Ni reduction of the chloride essentially as

described in Example 11 to produce the perfluorinated ether hydride.

5 Example 41. Preparation of $\text{CF}_3\text{O}(\text{CF}_2)_9\text{H}$ from $\text{CH}_3\text{O}(\text{CH}_2)_{10}\text{OAc}$

The starting material was prepared by monoalkylation of 1,10-decanediol with dimethyl sulfate, followed by acetylation with acetyl chloride. 10 This acetate was fluorinated, hydrolyzed, and decarboxylated essentially as in Example 1 to produce the perfluorinated ether hydride.

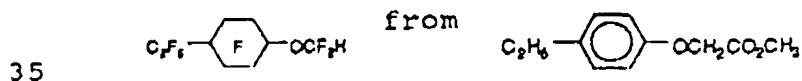
15 Example 42. Preparation of $\text{C}_9\text{F}_{19}\text{OCF}_2\text{H}$ from $\text{C}_9\text{H}_{19}\text{OC}_2\text{H}_4\text{OAc}$

The starting material was prepared by monoalkylation of ethylene glycol with n-nonyl bromide, followed by acetylation with acetyl chloride. This acetate was fluorinated, hydrolyzed, and decarboxylated 20 essentially as in Example 1 to produce the perfluorinated ether hydride (b.p. 155°C).

25 Example 43. Preparation of $(\text{iso-C}_3\text{F}_7)_2\text{CFOC}_2\text{F}_4\text{H}$ from $(\text{iso-C}_3\text{H}_7)_2\text{CHOC}_2\text{H}_4\text{CO}_2\text{CH}_3$

The starting material was prepared by Michael addition of 2,4-dimethyl-3-pentanol to acrylonitrile, followed by methanolysis to the methyl ester. This ester was fluorinated, hydrolyzed, and decarboxylated 30 essentially as in Example 1 to produce the perfluorinated ether hydride.

Example 44. Preparation of



The starting material was prepared by the alkylation of 4-ethylphenol with methyl chloroacetate. This ester was fluorinated, hydrolyzed, and decarboxylated essentially as in Example 1 to produce the perfluorinated ether hydride (b.p. 131°C).

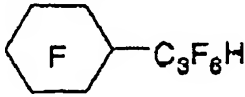
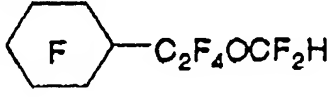

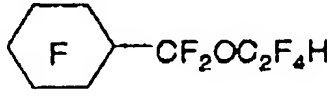
Example 45. Comparative Atmospheric Lifetimes and Boiling Points

The atmospheric lifetime of various sample compounds was calculated by the technique described in Y. Tang, Atmospheric Fate of Various Fluorocarbons, M. S. Thesis, Massachusetts Institute of Technology (1993). As shown in the table below, the atmospheric lifetime of an ether hydride compound having two or more carbon atoms between the ether oxygen atom and the terminal hydrogen atom is considerably shorter than the atmospheric lifetimes of ether hydride compounds having only one carbon atom between the ether oxygen atom and the terminal hydrogen atom. Because of the shorter atmospheric lifetimes of the compounds of the present invention, these compounds are more environmentally acceptable.

Compound	Atmospheric Lifetime (yrs)
$C_6F_{13}OC_2F_4OCF_2H$	>170
$C_4F_9OC_2F_4OCF_2H$	>170
$C_8F_{17}OCF_2CF_2H$	80

In addition, as shown in the table below, ether hydride compounds having two or more carbon atoms

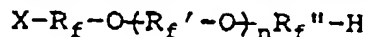
between the ether oxygen atom and the terminal hydrogen atom have lower boiling points than analogous non-ether compounds, and significantly lower boiling points than analogous ether hydride compounds having only one carbon atom between the ether oxygen atom and the terminal hydrogen atom. The unexpectedly low boiling points of compounds of the present invention render the compounds useful in processes involving temperature-sensitive substrates such as plastics. (For example, in vapor-phase cleaning, a substrate is rinsed in the condensing vapor of a boiling fluid, and in condensation heating, a substrate is heated by immersion in a boiling fluid.) In such applications, a lower-boiling fluid is preferred so as to avoid damage to the substrate. While it is known that boiling points can be reduced by selection of a compound having fewer carbon atoms, this may result in a boiling point reduction of 25°C or more, in addition to adversely affecting other properties such as solvency.

Compound	Boiling Point (°C)
$C_8F_{17}CF_2H$	136
$C_8F_{17}OCF_2H$	139
$C_7F_{15}OC_2F_4H$	131
	126
	134
	131
	125
$C_9F_{19}CF_2H$	154
$C_9F_{19}OCF_2H$	155
$C_8F_{17}OC_2F_4H$	148
$C_6F_{13}OC_4F_8H$	149
$C_5F_{11}OC_5F_{10}H$	150

Various modifications and alterations of this invention will become apparent to those skilled in the art without departing from the scope and spirit of this invention.

WHAT IS CLAIMED IS:

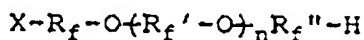
1. A normally liquid, omega-hydrofluoroalkyl
5 ether compound represented by the general formula:



wherein:

- H is a primary hydrogen atom;
X is a fluorine atom, a primary hydrogen atom, or
10 a primary chlorine atom;
n is an integer of 0 to 7; and
R_f, R_f', and R_f'' are independently selected from
the group consisting of linear or branched,
unsubstituted perfluoroalkylene groups;
15 linear or branched, perfluoroalkyl- or
perfluorocycloalkyl-substituted
perfluoroalkylene groups; and linear or
branched perfluoroalkylene groups substituted
with an ether oxygen-containing moiety;
20 with the proviso that when X is H or Cl, R_f has 1
to 18 chain carbon atoms, R_f' has 1 to 12 chain
carbon atoms, and R_f'' has 2 to 12 chain carbon
atoms;
and with the further proviso that when X is F, then R_f
25 has at least 4 chain carbon atoms, R_f' has 1 or
more chain carbon atoms, and R_f'' has 2 or more
chain carbon atoms.

2. A normally liquid, omega-hydrofluoroalkyl
30 ether compound represented by the general formula:



wherein:

- H is a primary hydrogen atom;
35 X is a fluorine atom, a primary hydrogen atom, or
a primary chlorine atom;
n is an integer of 0 to 7; and

R_f , R_f' , and R_f'' are independently selected from the group consisting of linear or branched, unsubstituted perfluoroalkylene groups; linear or branched, perfluoroalkyl- or perfluorocycloalkyl-substituted perfluoroalkylene groups; and linear or branched perfluoroalkylene groups substituted with an ether oxygen-containing moiety; with the proviso that when X is H or Cl, R_f has 1 to 18 chain carbon atoms and each of R_f' and R_f'' independently has 1 to 12 chain carbon atoms; and with the further proviso that when X is F, then R_f has at least 4 chain carbon atoms, and each of R_f' and R_f'' independently has 1 or more chain carbon atoms; and with the still further proviso that when n is zero, then R_f is a perfluorocycloalkyl-substituted perfluoroalkylene group.

3. A process for preparing an omega-hydrofluoroalkyl ether compound of claim 1, which comprises decarboxylating the corresponding precursor fluoroalkyl ether carboxylic acid, hydrolyzable derivative of said carboxylic acid, or hydrolyzable precursor to said carboxylic acid or said derivative, said decarboxylating being carried out by contacting said precursor carboxylic acid or ester with a solution of inorganic base in protic solvent and heating the resulting reaction mixture.

4. A method of displacing water from a surface which comprises contacting the surface with a liquid composition comprising the omega-hydrofluoroalkyl ether compound of Claim 1.

5. In a method of vapor phase soldering wherein a component to be soldered is immersed in or enveloped by a body of fluorinated liquid vapor to melt the solder, and the component is then withdrawn from the body of vapor, the improvement comprising using as the fluorinated liquid a composition comprising at least one omega-hydrofluoroalkyl ether compound of Claim 1.

6. A process for preparing a foamed plastic comprising the steps of:
admixing organic polyisocyanate and high molecular weight compound with at least 2 reactive hydrogen atoms in the presence of:

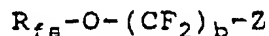
- 1) blowing agent mixture comprising at least one omega-hydrofluoroalkyl ether compound of claim 1;
- 2) a catalyst; and
- 3) a surfactant.

7. A method of removing a contaminant from an article comprising contacting said article with a composition comprising at least one omega-hydrofluoroalkylether compound of claim 1.

8. A method for the extinction of fires comprising applying to a fire a composition comprising at least one omega-hydrofluoroalkyl ether compound of claim 1.

9. A method for converting $-CF_2Cl$ groups to $-CF_2H$ groups comprising the step of contacting at least one compound containing at least one $-CF_2Cl$ group with hydrogen gas, said contacting being carried out at a temperature below about $200^\circ C$ in the presence of both a solution of base and a catalytic amount of at least one metal or supported metal, said metal being selected from the group consisting of nickel, palladium, and platinum.

10. A compound represented by the formula:



where

R_{fs} is a linear or branched perfluoroalkyl group
5 having from 1 to 18 carbon atoms;

"b" is an integer of at least 3; and

Z is selected from the group consisting of -COOH,

-COOM_{1/v}, -COONH₄, -COOR, -CH₂OH, -COF, -COCl,

-COR, -CONRR, -

CH₂NH₂, -CH₂NCO, -CN,

-CH₂OSO₂R, -

CH₂OCOR, -

CH₂OCOCR=CH₂,

-CONH(CH₂)_mSi(OR)₃,

and -

CH₂O(CH₂)_mSi(OR)₃,

where M is an

ammonium

radical or a metal atom having a valence "v"

of 1 to 4, each R is independently selected

from the group consisting of alkyl groups

having from 1 to 14 carbon atoms, fluoroalkyl

groups having from 1 to 14 carbon atoms, aryl

groups having from 6 to 10 ring-carbon atoms,

and heteroatom-containing alkyl groups having

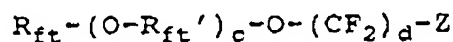
from 1 to 14 carbon atoms, fluoroalkyl groups

having from 1 to 14 carbon atoms, and aryl

groups having from 6 to 10 ring-carbon atoms,

and m is an integer of 1 to 11.

11. A compound represented by the formula:



wherein:

R_{ft} is a linear or branched perfluoroalkyl group
35 having from 1 to 18 carbon atoms;

R_{ft}' is a linear or branched perfluoroalkylene
group having from 1 to 11 carbon atoms;

c is an integer of at least 1;
d is an integer of at least 3; and
Z is selected from the group consisting of -COOH,
-COOM_{1/v}, -COONH₄, -COOR, -CH₂OH, -COF, -COCl,
-COR, -CONRR, -CH₂NH₂, -CH₂NCO, -CN,
-CH₂OSO₂R, -CH₂OCOR, -CH₂OCOCR=CH₂,
-CONH(CH₂)_mSi(OR)₃, and -CH₂O(CH₂)_mSi(OR)₃,

where M is an ammonium radical or a metal atom
having a valence "v" of 1 to 4, each R is
independently selected from the group consisting
of alkyl groups having from 1 to 14 carbon atoms,
fluoroalkyl groups having from 1 to 14 carbon
atoms, aryl groups having from 6 to 10 ring-carbon
atoms, and heteroatom-containing alkyl groups
having from 1 to 14 carbon atoms, fluoroalkyl
groups having from 1 to 14 carbon atoms, and aryl
groups having from 6 to 10 ring-carbon atoms, and
m is an integer of 1 to 11.

INTERNATIONAL SEARCH REPORT

 Internal Application No
 PCT/US 95/06110

 A. CLASSIFICATION OF SUBJECT MATTER
 IPC 6 C07C43/12 C07C59/135 C07C69/708 C07C41/18 C07C41/24
 B01D12/00 B23K35/38 C08G18/08 A62D1/00

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 C07C B01D B23K C08G A62D F26B

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP,A,0 261 501 (HOECHST) 30 March 1988 ---	1,2
X	JOURNAL OF FLUORINE CHEMISTRY, vol.10, no.1, July 1977, LAUSANNE CH pages 75 - 80 D. D. DENSON ET AL 'Synthesis of some perfluoroalkylether compounds' ---	1,2
X	JOURNAL OF HETEROCYCLIC CHEMISTRY, vol.10, no.6, December 1973, PROVO US pages 943 - 946 T. S. CROFT ET AL 'Fluoroalkyl-s-triazines' see paragraph bridging pages 944 and 945 --- -/-	1-3

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

* Special categories of cited documents :

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- *O* document referring to an oral disclosure, use, exhibition or other means
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T later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

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Date of the actual completion of the international search

14 September 1995

Date of mailing of the international search report

22.09.95

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Wright, M

INTERNATIONAL SEARCH REPORT

Internat. Application No.

PCT/US 95/06110

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	TETRAHEDRON LETTERS., vol.30, no.48, 1989, OXFORD GB pages 6717 - 6720 CHANG-MING HU ET AL 'Mechanistic study on the photooxidation of perhalofluoroalkyl sulfinates' ---	1,2
X	US,A,3 706 773 (L. G. ANELLO ET AL) 19 December 1972 see examples 7-10,12-14,16-18 ---	10
X	US,A,3 674 800 (R. F. SWEENEY ET AL) 4 July 1972 see examples 1,5 ---	10
X	US,A,3 766 274 (L. G. ANELLO ET AL) 16 October 1973 see examples 12,13 ---	10
X	JOURNAL OF ORGANIC CHEMISTRY OF THE USSR., vol.15, no.7, July 1979, NEW YORK US pages 1264 - 1271 V. V. BERENBLIT ET AL 'Synthesis and some transformations of derivatives of perfluoroalkoxycarboxylic acids' see pages 1266 and 1268, compound nos.VII-X,XXXIII ---	10
A	FR,A,1 373 014 (E. I. DU PONT DE NEMOURS) 17 August 1964 see page 5, column 2; claims -----	5

INTERNATIONAL SEARCH REPORT

International Application No
PCT/US 95/06110

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP-A-0261501	30-03-88	DE-A- 3631560 DE-A- 3785121	14-04-88 06-05-93
US-A-3706773	19-12-72	BE-A- 714159 BE-A- 714160 BE-A- 714161 BE-A- 714162 DE-A- 1768286 DE-A- 1768287 DE-A- 1768288 DE-A- 1770262 FR-A- 1574801 FR-A- 1578003 FR-A- 1599703 FR-A- 1580225 GB-A- 1224771 GB-A- 1224773 GB-A- 1224772 GB-A- 1224774 NL-A- 6805860 NL-A- 6805861 NL-A- 6805862 NL-A- 6805864 US-A- 3514487 US-A- 3577465 US-A- 3647887 US-A- 3697564 US-A- 3781370	16-09-68 16-09-68 16-09-68 16-09-68 28-10-71 18-11-71 18-11-71 30-03-72 18-07-69 14-08-69 20-07-70 05-09-69 10-03-71 10-03-71 10-03-71 10-03-71 28-10-68 28-10-68 28-10-68 28-10-68 26-05-70 04-05-71 07-03-72 10-10-72 25-12-73
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US-A-3766274	16-10-73	DE-A- 2051551 GB-A- 1308794	29-04-71 07-03-73
FR-A-1373014	04-01-65	NONE	

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US 95/06110

Box I Observations where certain claims were found unsearchable (Continuation of item 1 of first sheet)

This international search report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. ☐ Claims Nos.:
because they relate to subject matter not required to be searched by this Authority, namely:

2. ☒ Claims Nos.:
because they relate to parts of the international application that do not comply with the prescribed requirements to such an extent that no meaningful international search can be carried out, specifically:
Claim 9 : the search has been limited to the preparation of ethers
Claims 10 and 11 : the documentary search has been limited to acids,
salts and esters (Rule 33.3 PCT)

3. ☐ Claims Nos.:
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

Box II Observations where unity of invention is lacking (Continuation of item 2 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

1. ☐ As all required additional search fees were timely paid by the applicant, this international search report covers all searchable claims.

2. ☐ As all searchable claims could be searched without effort justifying an additional fee, this Authority did not invite payment of any additional fee.

3. ☐ As only some of the required additional search fees were timely paid by the applicant, this international search report covers only those claims for which fees were paid, specifically claims Nos.:

4. ☐ No required additional search fees were timely paid by the applicant. Consequently, this international search report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

- ☐ The additional search fees were accompanied by the applicant's protest.
- ☐ No protest accompanied the payment of additional search fees.

A Status Report on the Development of HFC Blowing Agent for Rigid Polyurethane Foams

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ABSTRACT

The field of blowing agents for rigid polyurethane foams has seen a great deal of upheaval in the recent years. Starting with a situation where CFC-11 was the only blowing agent in use, various alternatives have emerged, among which HCFC-141b has become the product of choice when high thermal and fire performance are to be achieved. It is however already time to think again about a new generation of blowing agents that would have no impact on the ozone layer.

This paper provides an update on the actions undertaken by Elf Atochem (leading supplier to the rigid polyurethane foam market with FORANE® 141b) to prepare the future by developing zero ODP/high performance alternatives.

This will include the presentation of our latest laboratory results regarding the evaluation of several liquid and gaseous HFCs as blowing agents in rigid polyurethane foams such as HFC-134a, HFC-245fa, HFC-365mfc, HFC-245eb, HFC-245ca, HFC-236ea. The results of the evaluation of HFC-365mfc and HFC-134a by foam producers or raw material suppliers in different applications will also be presented.

CURRENT REGULATIONS ON OZONE DEPLETING SUBSTANCES

The discovery in the 70-80's that some molecules could have a very severe effect on the ozone layer has prompted the international community to enact stringent regulations on their production and/or uses.

From its enactment in 1987 and through its subsequent reviews, the now well-known Montreal Protocol has been the regulatory arm chosen by the international community for the control of substances that deplete the ozone layer. This international agreement provides a phase-out schedule that will differ depending on the type of substance con-

sidered and the group the country belongs to ("developed countries" and "article 5 countries"). For the products used as blowing agents in polyurethane foams, according to the Montreal Protocol, the production of CFCs in developed countries was forbidden on 1/1/1996. The "article 5 countries" will have until 2010 to end their production. HCFCs have also been included and have an elimination schedule of 2020 (gradual reduction from 2004 to 2020) for the developed countries and 2040 for the article 5 countries (gradual reduction from 2016 to 2040).

Within the frame of this protocol, some countries, such as the United States and the European Union, have decided to take further steps towards the total phase-out of all ozone depleting substances. The United States has decided, for example, HCFC-141b will have to be phased-out in 2003. Meanwhile, the European Union has phased-out CFCs on 1/1/1995. It has further placed restrictions on the uses of the HCFCs, limiting them to non-emissive applications, such as rigid insulating foams. In the European Union, HCFCs will be eliminated in 2014. This last regulation can even lead to some unexpected results since in the European Union, using HCFCs for flexible or integral skin polyurethane foams is therefore forbidden, but the use of CFCs is not.

Figure 1 summarizes these regulations. It is necessary to keep the whole picture in mind to understand the current developments in the field of blowing agents for rigid polyurethane foams.

HCFC-141b: A SUCCESSFUL BLOWING AGENT

Since CFCs (CFC-11 in the case of rigid polyurethane foams) have by far the greatest Ozone Depletion Potential (as shown in Figure 2), the most effective way to help restore the ozone layer is to stop using them as soon as possible.

In developed countries, the phase-out of CFC-11 was made possible by the availability of substitutes (HCFC-

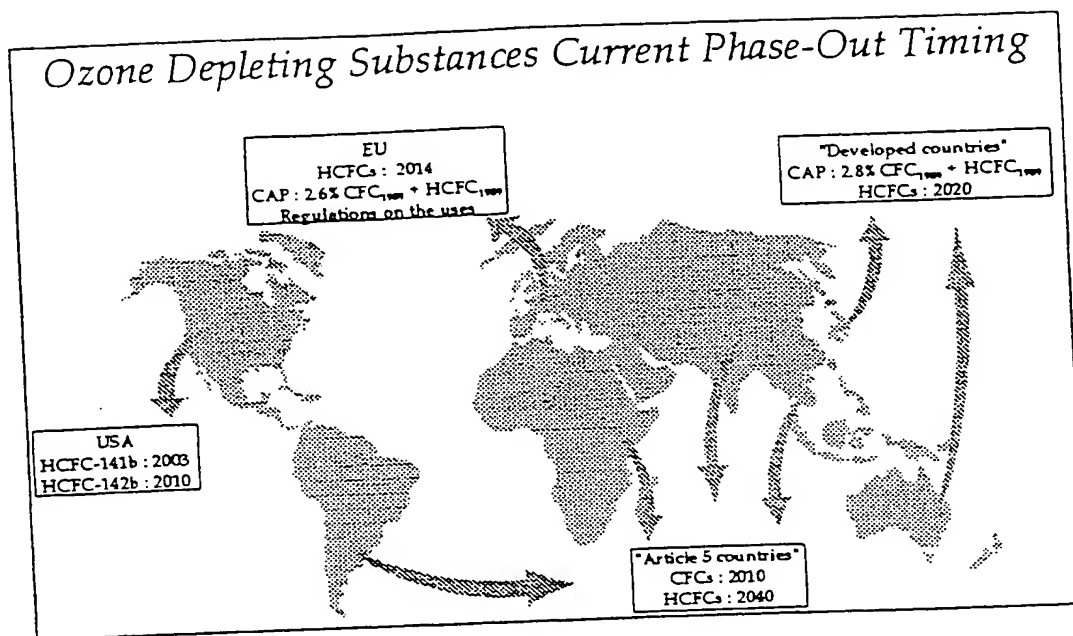


Figure 1.

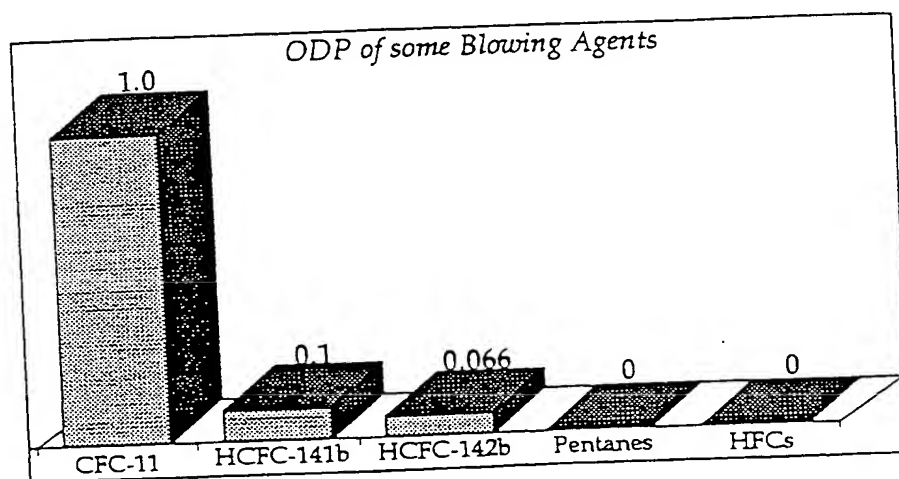


Figure 2.

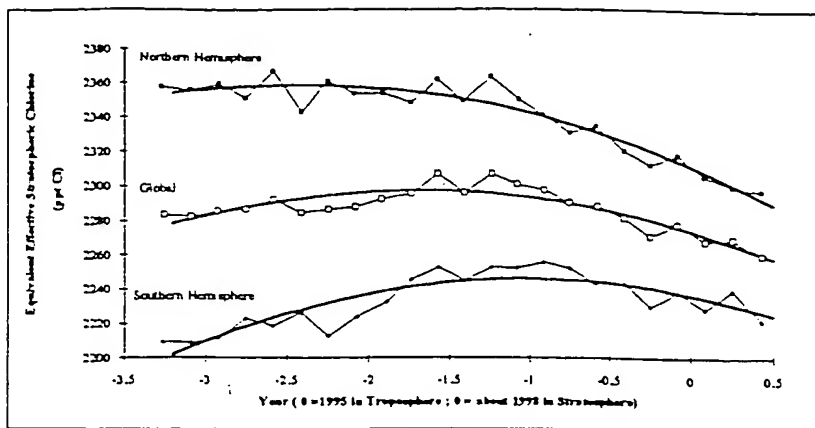


Figure 3.

141b in the US, HCFC-141b and hydrocarbons in the European Union). The phase-out of CFC-11 in foams, along with those made by the other application sectors (aerosols, refrigeration, fire extinguishers . . .), which were using ozone depleting substances, has already proven to be successful. The last measurements of the atmospheric chlorine loading [1] clearly show that we are on the right track. Indeed, as shown in Figure 3, the equivalent effective chlorine content in the northern troposphere (lower layer of the atmosphere) has been decreasing for a few years. Since chemical compounds in the troposphere need about 3 years to migrate to the stratosphere (upper layer of the atmosphere—where the ozone layer is), we can expect to observe the same encouraging findings in the stratosphere before the end of the century; hence, the ozone layer should begin its recovery process very soon despite the fact that recovery will not be completed until several decades from now.

Therefore, the Montreal Protocol is already a success!

Now, the most important issue is to help "article 5" countries to stop using CFCs as quick as possible. HCFCs, especially HCFC-141b, are the easiest and most efficient way to accomplish this goal.

There is concern that the European Union wants to expedite the total elimination of HCFC-141b. Such decision would not only have no meaningful effect on the recovery of the ozone layer, but also would send the wrong message to "article 5" countries: encourage them to continue using CFC-11 until 2010 rather than to switch to a possible threatened blowing agent like HCFC-141b or to hydrocarbons, with the large investments they require to be handled safely.

In spite of its small effect on the ozone layer, HCFC-141b has proven to be, and still is a necessary step towards the total elimination of CFC-11.

ZERO ODP GASEOUS BLOWING AGENTS

Even if HCFC-141b is still definitely necessary to the rigid polyurethane foam industry, the development of zero ODP substitutes has been ongoing for several years.

Among different candidates, HFCs have already appeared as the most promising candidates. They are efficient blowing agents, safe to handle and have a minimized overall impact on the environment [2]: zero Ozone Depletion Potential (ODP), low direct Global Warming Potential, very good energy efficiency and therefore a reduced indirect impact on the greenhouse effect (Total Equivalent

Warming Impact concept), very low Photochemical Ozone Creation Potential (the HFCs do not participate in the creation of atmospheric pollutant and therefore are not usually classified as Volatile Organic Compounds).

Among the HFCs, it is possible to distinguish between the gaseous and the liquid ones. Gases are not the preferred options for the expansion of polyurethane foams since they require specific handling procedures, pressure vessels, and may cause some frothing when used. However, since some gaseous HFCs are already available, they can be an appropriate answer for those who want to use immediately a zero ODP, non flammable blowing agent. HFC-134a, initially developed in refrigeration, is the first obvious choice to look at. Its main properties are summarized in Table 1.

As already reported elsewhere [3], the main drawback of HFC-134a is its limited solubility in most polyols for rigid polyurethane foams. To establish that aspect, its solubility has been determined experimentally in different polyols. The polyols tested are listed in Table 2.

The experimental set-up is as follows:

- a known mass of the polyol, pure or blended, is introduced inside an aerosol can
- a known amount of HFC-134a is then added
- the can is vigorously stirred for half an hour
- once the temperature of the can is stabilized at 25°C, the partial pressure of the blowing agent inside the can is measured

When applying this procedure, a plot of the blowing agent partial pressure versus the amount of HFC-134a solubilized inside the polyol can be obtained. Figure 4 displays the results obtained for four different polyols with

Table 1.

	HCFC-141b	HFC-134a
Formula	CH ₃ CFCl ₂	CF ₃ CFH ₂
MW (g/mol)	117	102
B.P. (°C)	32.1	-26.4
d (g/cc) @20°C	1.24	1.22
λ _{gas} @25°C (mW/mK)	9.7	13.5
Flash point	None	None
Flame lim. (% vol.)	5.8-17.7	None
ODP	0.10	0
GWP _{100yr}	530	1300
Toxicity	Good	Good

Table 2.

	Supplier	Polyether Type
3750	ARCO	Aromatic amine based
3502	ARCO	Blend of polyethers: sorbitol based and glycerine based
3544	ARCO	Sorbitol based
3528	ARCO	Sorbitol based

HFC-134a. At a pressure of one bar, the amount of HFC-134a dissolved in polyols can be obtained; this data will indeed correspond roughly to the amount that it will be possible to incorporate in a polyol system and still keep the blend polyol/HFC-134a in a non pressurized container.

Except for the aromatic amine based polyether polyol (A3750), the amount of HFC-134a that can be solubilized in the other polyols, for a partial pressure of HFC-134a equal to one bar, is quite constant and is around 0.035 mole/100 g of polyol. This means therefore that around 3.5 parts of HFC-134a can be included in 100 parts of these polyols.

HFC-134a industrial scale trials have already been conducted in various locations throughout the European Union. Because of the limited solubility of HFC-134a in the standard polyols, most of the systems used so far included only low levels of HFC-134a (2-3 parts of HFC-134a in 100 parts of polyol) and were therefore largely water co-blown. Table 3 presents typical properties that can be achieved with this system.

As noted, results are quite satisfactory and could be improved by the use of higher HFC-134a levels perhaps by the direct injection of HFC-134a at the injection head.

ZERO ODP "LIQUID" BLOWING AGENTS—FOAMING TESTS

As we have shown earlier, HFC-134a can be a solution for those who want to use a non-flammable, zero ODP blowing agent immediately. There is, however still a need for a true "liquid" HFC substitute to HCFC-141b which will be able to replace it in all its applications. Several candidates have already emerged. A list of those candidates along with their principal physical properties is shown in Table 4. Among them are the three isomers of pentafluoropropane, HFC-365mfc, HFC-236ea and HFE-236 (which is in fact a commercially available molecule used as anaesthetic).

All these molecules were tested as blowing agents in the

same conditions. The formulations used are reported in Table 5. For comparison purposes, the same base polyol/catalyst/surfactant was used for the six HFCs and HCFC-141b was used as a baseline. Moreover, all the reactants were mixed at 5°C (hand made foams) in order to accommodate HFC-236ea, the blowing agent with the lowest boiling point in this set. An additional group of HCFC-141b foams was also made with all the chemical reactant kept at 20°C. This allows us to evaluate the influence of initial temperature.

Figures 5, 6 and 7 show the main characteristics of the foams obtained with the various systems.

From these results, the following trends can be outlined:

- The three pentafluoropropane isomers, namely HFC-245ca, HFC-245eb, HFC-245fa, yield good properties in all the areas surveyed: thermal conductivity (both initial and aged), compression set and open cell content.
- For this non-optimized formulation, HFC-365mfc stands just behind the previous blowing agents. Using a more suitable surfactant (see hereafter) would undoubtedly improve the results.
- On the other hand, the relatively low boiling point of HFC-236ea appears to play against this blowing agent since controlling the foam structure seems to be more difficult than with the other blowing agents (presence of holes and chimneys).
- Even if all the characteristics of HFE-236 seem to favour this compound as a blowing agent for rigid polyurethane foams (boiling point, gaseous phase thermal conductivity), the actual foaming tests reveal that its solubility in the polyurethane matrix is too high to have good foam performance.
- At last, the comparison of the polyurethane foams blown with HCFC-141b at 20°C and at 5°C shows that the later is close to the former except that the thermal conductivity is 1 mW/mK higher.

ZERO ODP "LIQUID" BLOWING AGENTS—COMPARATIVE DIFFUSION STUDY

Being able to assess the long term thermal properties of the polyurethane foams blown with these liquid HFCs is also a major issue in the substitution of HCFC-141b. One way to deal with this question is to observe the foam's thermal conductivity ageing. Such data have already been presented in Figure 5. Another possibility is to monitor the gas phase composition of a polyurethane foam to analyse how

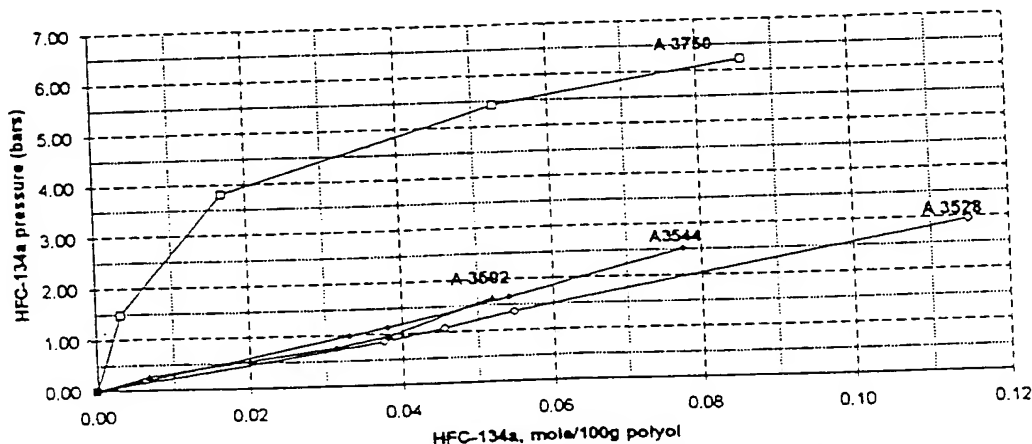


Figure 4. HFC-134a volatility vapor pressure of polyol mixtures @ 25°C.

Table 3. Industrial trial for HFC-134a/H₂O systems in metal faced panels.

Thickness (mm)	70
Core density (g/l)	41.5
Compression set @ 10% (kPa)	140
Open cell content (%)*	2.3
Thermal conductivity @ 24°C (mW/mK)	
Initial	22.4
After ageing (7w @ rt + 9w @ 70°C)	23.1
Core cell gas composition: volume fractions	
Air/CO ₂ /HFC-134a	
Initial	1.6/94.7/3.7
After ageing (7w @ rt + 9w @ 70°C)	23.8/68.4/7.8
Dimensional stability (% vol. change)	
5d @ 90°C	-4.5
5d @ -18°C	-0.5

*ASTM D 2856—Procedure B with an initial pressure step of 0.75 psi.

Table 4.

	HCFC-141b	HFC-236ea	HFC-365mfc	HFC-245ca	HFC-245eb	HFC-245fa	HFE-236
Formula	CH ₃ CFCl ₂	CF ₃ CHFCF ₂ H	CF ₃ CH ₂ CF ₂ CH ₃	CFH ₂ CF ₂ CF ₂ H	CF ₃ CHFCH ₂ F	CF ₃ CH ₂ CF ₂ H	CF ₃ CHFOCF ₂ H
MW (g/mol)	117	152	148	134	134	134	168
B.P. (°C)	32.1	6.5	40.0	25.4	22.7	15.3	22.8
d (g/cc) @20°C	1.24	1.40	1.25	1.40	1.38	1.32	1.45
λ _{gas} @25°C (mW/m.K)	9.7	11.1*	10.6*	10.8*	11.0*	11.2*	11.6*
Flash point	None	None	-25°C	None	None	None	None
Flame lim. (% vol.)	5.8-17.7	None	3.5-9.0	8.3-12.8	9.6-10.7	None	N.A.
ODP	0.10	0	0	0	0	0	0
HGWP	0.14	0.35	0.21	0.15	0.12	0.24	N.A.
Toxicity	Good	Pending	Pending	Pending	Pending	Pending	Good

*Calculated data.

Table 5.

	Chemical Temperature (A/B)							
	20°C				5°C			
PU 1832 (BAYER system)	100	100	100	100	100	100	100	100
44v20 (iso index 105)	163	163	163	163	163	163	163	163
HCFC-141b	21	21						
HFC-245ca			24.4					
HFC-245eb				24.4				
HFC-245fa					24.4			
HFC-236ea						27.7		
HFC-365mfc							27	
HFE 236								30.6
Core density (g/l)	35.8	35.4	36.5	36.5	34.1	33.0	35.9	32.6

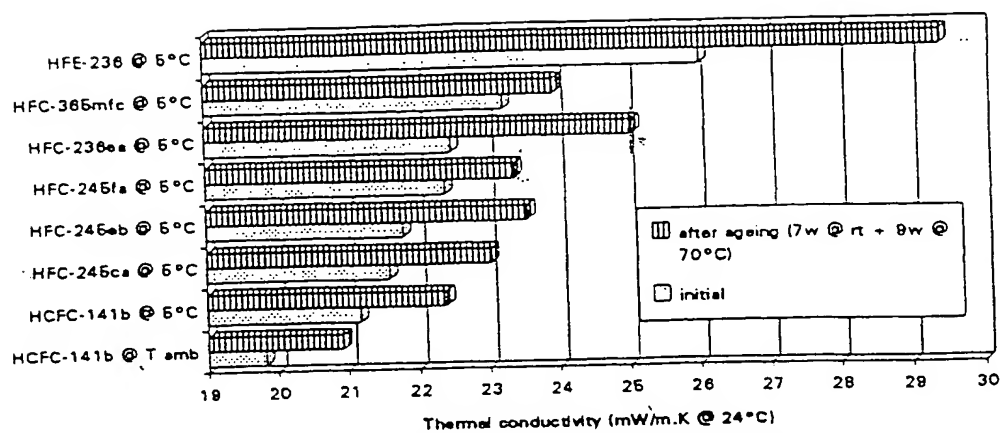


Figure 5. Blowing agent's effect on the thermal conductivity.

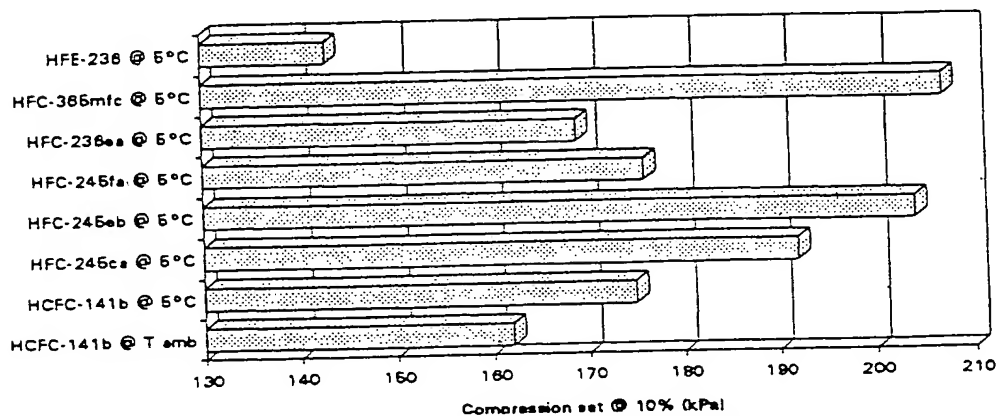


Figure 6. Blowing agent's effect on the compression set.

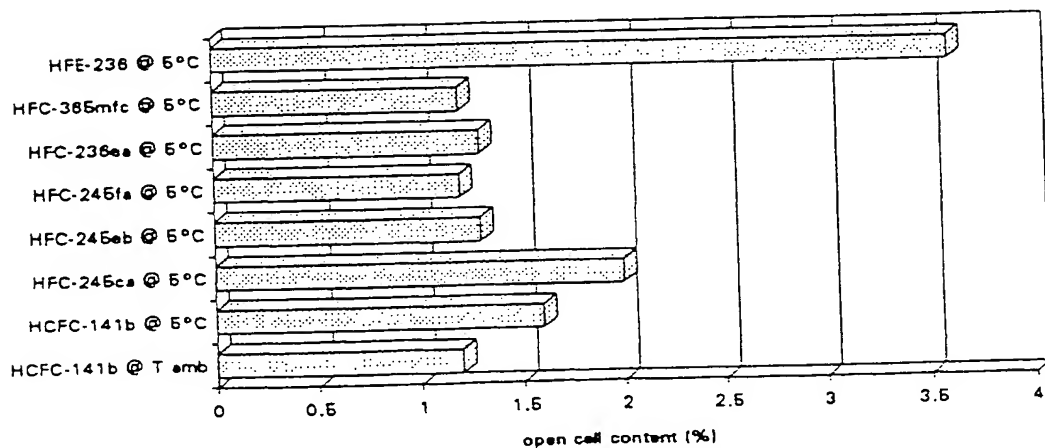


Figure 7. Blowing agent's effect on the open cell content. (ASTM D2856—Procedure B with an initial pressure step of 0.75 psi.)

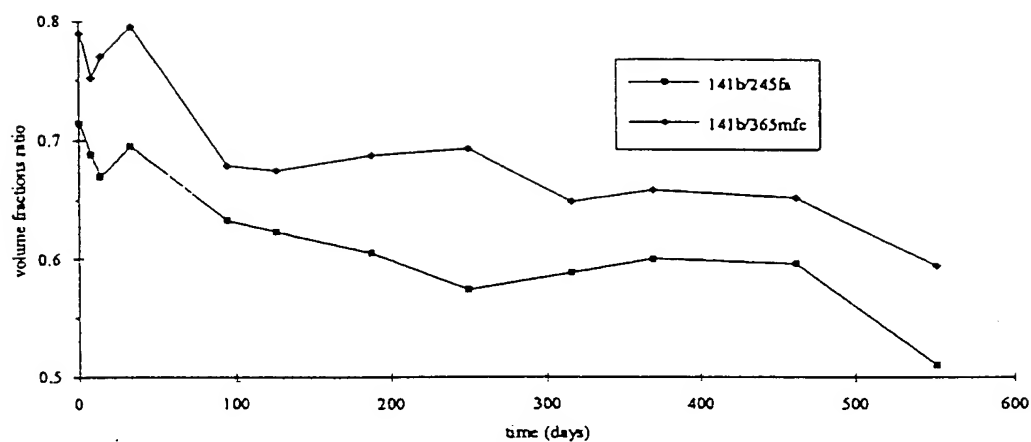


Figure 8. Comparative diffusion HCFC-141b/HFC-245fa/HFC-365mfc @ rt.

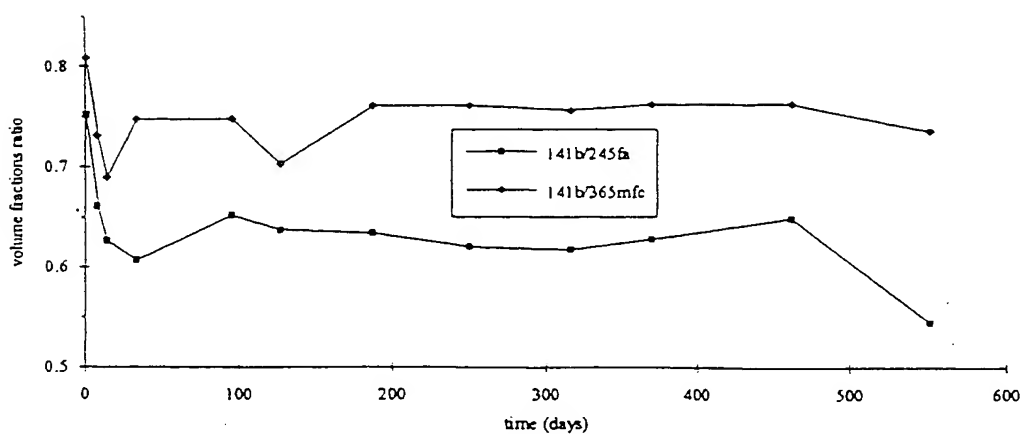


Figure 9. Comparative diffusion HCFC-141b/HFC-245fa/HFC-365mfc @ 70°C.

Table 6.

	Initial Volume Fractions (%)				
	Air	CO ₂	HFC-245fa	HCFC-141b	HFC-365mfc
Sample 1 (to be aged @ rt)	1.5	38.5	22.4	16.8	20.8
Sample 2 (to be aged @ 70°C)	1.3	40.8	22.1	15.8	20.0

Table 7. Surfactants evaluation for HFC-365mfc based systems by OSI Specialties S.A.

Polyol	100
Isocyanate	175
Water	1.8
DMCHA	1.5
Silicone	0.3 to 1.5
HFC-365mfc	31

Table 8. Surfactant evaluation for HFC-365mfc based systems by OSI Specialties S.A.

	HFC-365mfc Flowability* (cm/g)	HFC-365mfc Initial λ -value (mW/m.K @ 24°C)**	HCFC-141b (reference) Initial λ -value (mW/m.K @ 24°C)**
NIAX® L-6900	0.95	22.7	22.1
NIAX® L-6908	0.93	23.0	21.7
NIAX® SR-321	0.72	23.1	/

*0.3 part of silicone

**1.5 part of silicone—Moulded foam: 34 g/l—free rise density: 25.0 g/l.

Table 9. HFC-365mfc evaluation by Industrial Urethanes (system supplier—South Africa).

Polyol blend	79.8
Isocyanate (index 110)	128
Water	1.8
HFC-365mfc	18.4

Table 10. HFC-365mfc evaluation by Industrial Urethanes (system supplier—South Africa).

Cream time (hand mix)	15 seconds
String time	67 seconds
Tack free time	103 seconds
Free rise density	25.6 kg/m ³
Initial k -value @ 10°C	22.1 mW/mK
Compression strength 10%	247.4 kPa (parallel to rise)
Samples from Brett Mould (foam core density 31.8 kg/m ³)	106.0 kPa (perpendicular to rise)

fast these new blowing agents will leave the polyurethane foam (the long term thermal efficiency of a closed cell rigid polyurethane foam depends very much on the blowing agent diffusion out of the foam).

To that intent, a method has been developed to periodically analyse the content of the gas phase of a cellular polymer. The experimental set-up is relatively easy to implement: under a helium flow, a syringe needle is pushed inside the foam; then, the content of the broken cells is pulled into the syringe's body. Keeping the needle inside the helium stream, the content of the syringe is injected in a gas chromatogram and is analyzed for air, CO₂ and the various blowing agents. The helium flow is necessary in order to prevent any outside air from polluting the original cell content (which could happen since pressure within the cells might be below atmospheric pressure). With this method, it is thus possible to determine the volume fractions of all the components present in the gas phase of a foam. Moreover, since it is non destructive (only a few cells are broken for each determination), it is possible to follow the evolution of the cell gas composition with time.

The first step to study the comparative diffusion of some of the most likely substitutes has been to blow two polyurethane foams with a mixture HCFC-141b/HFC-245fa/HFC-365mfc. The initial cell gas contents analyzed according to the above procedure for the two polyurethane samples are presented in Table 6.

Subsequent to this analysis, one sample has been aged at room temperature and the other at 70°C.

In this case, we have followed the ratio of the volume fraction of HCFC-141b over HFC-245fa or over HFC-365mfc. Indeed, the volume fraction of HCFC-141b over the volume fraction of HFC-245fa is simply the ratio of the amount of HCFC-141b present in the gaseous phase over the amount of HFC-245fa in this same gaseous phase; the total volume, which would depend on the amount of air or CO₂, is no longer involved.

If we take a closer look at the amount of a physical blowing agent (HCFC-141b, HFC-245fa or HFC-365mfc) in the gaseous phase, it can only decrease through several mechanisms: diffusion of the blowing agent outside the foam, dissolution inside the polymer matrix or condensation within the cells. The only possibility to observe an increasing amount of blowing agent would have to come from a desorption from the matrix into the gas phase. Under usual circumstances, this is unlikely to happen.

So, for example, if the ratio of the volume fraction of HCFC-141b over HFC-245fa is increasing, it necessarily means that the amount of HFC-245fa in the gas phase has decreased faster than the amount of HCFC-141b. On the other hand, if this ratio has decreased, it necessarily means that the amount of HCFC-141b has decreased faster than the amount of HFC-245fa.

Figure 8 represents the evolution of the ratio of the volume fractions of HCFC-141b over HFC-245fa and of HCFC-141b over HFC-365mfc for the polyurethane sample aged at room temperature. Figure 9 gives the same information for the foam aged at 70°C.

Let's analyse first the evolution at room temperature. In both cases, the ratio of the volume fractions of HCFC-141b over HFC-245fa or over HFC-365mfc is clearly decreasing with time, meaning that HCFC-141b is disappearing faster from the foam's gas phase than HFC-245fa or HFC-365mfc. Therefore, we can expect a slower ageing (better aged thermal conductivity) for a polyurethane foam blown with either of these two HFCs than with HCFC-141b. As for the evolution at 70°C, the two ratios have again a marked de-

creasing trend at the beginning but a more constant aspect afterward. From this comparative diffusion study, we can say that long term thermal ageing of the foam should be at least as good with either HFC-245fa or HFC-365mfc as with HCFC-141b itself.

Our results show that efficient zero ODP substitutes for HCFC-141b do exist. Among them, HFC-245fa and HFC-365mfc appear to be the most promising when not only their performance as blowing agents for rigid polyurethane foams but also their ease of manufacture is considered. Moreover, HFC-365mfc has one more advantage over the other candidates since it can be a HCFC-141b substitute not only as a blowing agent but also as a solvent, which would expand its potential market, and therefore favourably impact its economics.

ZERO ODP "LIQUID" BLOWING AGENTS—HFC-365mfc CUSTOMERS' EVALUATIONS

All these facts have prompted Elf Atochem to pursue the evaluation of HFC-365mfc in collaboration with several PUR foam producers or raw material suppliers. As an example, OSi Specialties S.A. has been working on the identification of silicone surfactant especially suitable for HFC-365mfc based systems. Table 7 lists a typical laminated board formulation that was used by OSi Specialties.

Table 8 provides some of the results and shows that NIAX® L-6900 and L-6908 seem to give excellent performance with HFC-365mfc systems in terms of both foam flowability and thermal conductivity.

Finally, HFC-365mfc has also been evaluated by Industrial Urethanes (system supplier—South Africa) in the formulation described in Table 9 and has yielded the results presented in Table 10.

These results are again very encouraging and clearly demonstrate the potential of HFC-365mfc as a blowing agent for rigid polyurethane foams.

CONCLUSION

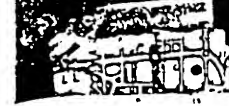
Since many factors are necessary to qualify a candidate as an acceptable blowing agent (environmental impact, toxicity, flammability, foam performance) and since several other factors must be considered for a candidate to be selected as a commercial product (availability of feedstock, ease of manufacture, cost, market size), the objective for Elf Atochem is to be able to determine the best option within the next years, the following years being possibly dedicated to long term toxicological studies, pilot and plant construction and reformulation.

ACKNOWLEDGEMENTS

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BIOGRAPHIES

Jean-Denis Roux

Jean-Denis Roux is an Engineer with Elf Atochem S.A. based at their Application Laboratory. Since 1992, he has been in charge of the FORANE range development/technical service in the area of blowing agents for cellular plastics.

Arnaud Albouy

Arnaud Albouy graduated in 1993 from the Ecole Supérieure de Physique et Chimie Industrielles de la Ville de Paris. He joined Elf Atochem in 1995. His current responsibilities are the development of blowing agents for plastic foams.

Denis Mouton

Denis Mouton has been working on plastic foams for the last 25 years. He joined Elf Atochem in 1989 where he is Business Manager, in charge of the FORANE range (blowing agents/solvents).

Jinhuang Wu

Jinhuang Wu is a scientist at Elf Atochem North America's Technological Centre in King of Prussia. He received his PhD in Macromolecular Science and Engineering from the University of Michigan and joined Elf Atochem in March 1997.